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| Introduction to Concentration UnitsCλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues, Weds., &Thurs after school!After-hours question? Email me at home: eagenest@madison.k12.wi.us |  | Name\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_ |

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|  |  | 1. Find the molarity of 350.mL of a solution that contains 0.0049 moles of NaCl.
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|  |  | 1. Find the molarity of 350.mL of a solution that contains 0.0049 grams of NaCl.
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|  |   | 1. How many moles of CH3Br are dissolved

in 40. mL of 2 M solution? |

1. 
2. 
3. 
4. 
5. 
6. Circle the metallic element in each.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| *Circle any element that is a metal* | This substance is… |  | *Circle any element that is a metal* | This substance is… |
| Al(NO3)3 | ionic / molecular |  | N2H4O | ionic / molecular |

1. How many pieces will this fall apart into if made into an aqueous solution? (circle your choice)

|  |  |  |  |
| --- | --- | --- | --- |
| (NH4)2CO3 | 1? 2? 3? 4? 5?  | Al(NO3)3 | 1? 2? 3? 4? 5?  |
| K2CO3 | 1? 2? 3? 4? 5?  | Li2S | 1? 2? 3? 4? 5?  |

1. Write your answer with a coefficient in front of each particle*. For example, instead of writing Br- Br- Br- it is more common and easier to write a coefficient: 3Br-*
	1. Li2S(s) dissolving. Use a coeffiecient where appropriate.

\_\_\_\_\_\_\_\_ --> \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_

* 1. Al(NO3)3 (s) dissolving. Use a coeffiecient where appropriate.

\_\_\_\_\_\_\_\_ --> \_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_