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| Wrong Lewis Dot Molecule StructuresCλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues,&Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_ |

1. Fill in the table

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| Circle atoms that are stable, cross out atoms that are unstable | Either write ‘stable’ or redraw the same letters in the same arrangement but with any number of e- that will make the atoms stable | Circle atoms that are stable, cross out atoms that are unstable | Either write ‘stable’ or redraw the same letters in the same arrangement but with any number of e- that will make the atoms stable |
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1. Rewrite the following Noble Gas Abbreviations in the longer version of electron configuration (1s2 2s2 etcetera)
	1. [Ne]3s23p2
	2. [He]2s22p1
	3. Write the Lewis Dot symbol for each of the two atoms above:
2. Write a balanced equation for N3- anion losing three electrons:

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1. Write a balanced equation for Na+ ion gaining one electron:

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