|  |  |  |
| --- | --- | --- |
| Wrong Lewis Dot Molecule Structures  CλeMis+ry: http://genest.weebly.com  Stop in for help every day at lunch and Tues,&Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_  Period\_\_\_\_\_\_\_\_ |

1. Fill in the table

|  |  |  |  |
| --- | --- | --- | --- |
| Circle atoms that are stable, cross out atoms that are unstable | Either write ‘stable’ or redraw the same letters in the same arrangement but with any number of e- that will make the atoms stable | Circle atoms that are stable, cross out atoms that are unstable | Either write ‘stable’ or redraw the same letters in the same arrangement but with any number of e- that will make the atoms stable |
|  |  |  |  |

1. Rewrite the following Noble Gas Abbreviations in the longer version of electron configuration (1s2 2s2 etcetera)
   1. [Ne]3s23p2
   2. [He]2s22p1
   3. Write the Lewis Dot symbol for each of the two atoms above:
2. Write a balanced equation for N3- anion losing three electrons:

\_\_\_\_ \_\_\_\_

1. Write a balanced equation for Na+ ion gaining one electron:

\_\_\_\_ \_\_\_\_