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| Draw polar bondsCλeMis+ry: Final exam study info at http://genest.weebly.com Start making your cheat sheet for the final exam. | look up Mach Illusion on the internet | Name\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_ |

1. According to your electronegativity table, what is the exact electronegativity of each element below?
	1. beryllium
	2. phosphorous
	3. calcium
	4. fluorine
2. Calculate the number of valence e- in each of the following:
3. SO42-
4. NH4+
5. CH3O2+
6. ClO3–
7. Next to each atom, write the electronegativity value from your table. Then draw an arrow to show the direction of polarity of the bond. Finally, write the symbols for partial positive and negative.δ+ and δ-.



1. Why does your table not have electronegativity numbers for the nobel gases?
2. Next to each, write the name of the geometric shape based on the central atom. **Assume each line is a connection between two atoms.**

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1. Draw a stable Lewis Dot structure for each molecule AND THEN WRITE the name of the molecules geometry, based on the central atom:

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| 1. NF3

What’s the name of this geometry? | 1. SO42-

What’s the name of this geometry? |
| NH4+What’s the name of this geometry? | ClO3–that’s chlorine with three oxygen atomsWhat’s the name of this geometry? |

1. How many total valence e- are in each of the following?
	1. I3-
	2. CO32-
	3. PO43-
	4. NO3-
2. Which element in Group 2 has the largest radius?
3. Which element in Period 2 has the largest radius?
4. Draw a stable Lewis Dot structure for each molecule:

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| I3-What’s the name of this geometry? | CO32-What’s the name of this geometry? |
| NO3-What’s the name of this geometry? | SCl2What’s the name of this geometry? |