

Ionization Energy for removing all the electrons from BORON, one by one

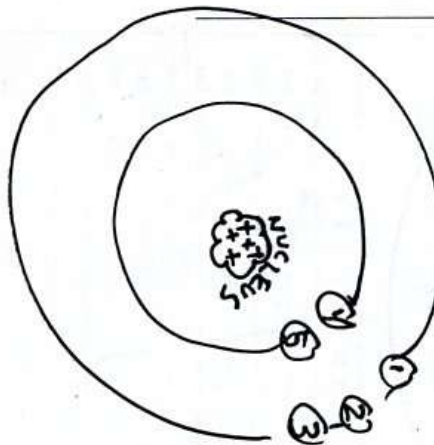
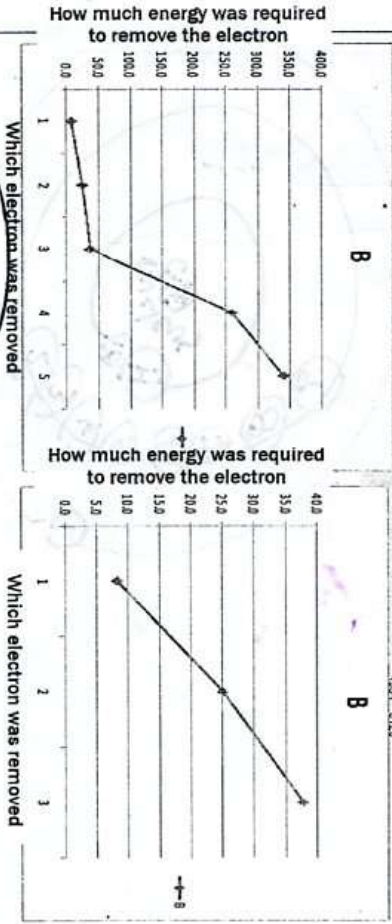
Below are graphs of the ionization energies for BORON.

Two definitions to memorize:

Electron Affinity is defined as the energy required to *add* an electron to an atom.

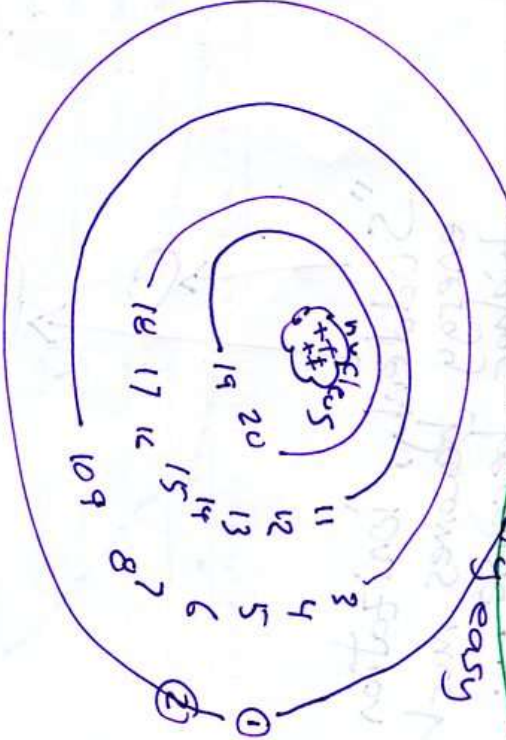
Ionization energy is defined as the energy required to *remove* an electron from an atom. It is not the energy for adding an electron.

The graphs here are for ionization energy only.



Handwritten notes:
 for boron graph
 energy for 2nd shell
 1st shell (2 electrons)
 2nd shell (3 electrons)

Ionization Energy for removing all the electrons from Calcium, one by one

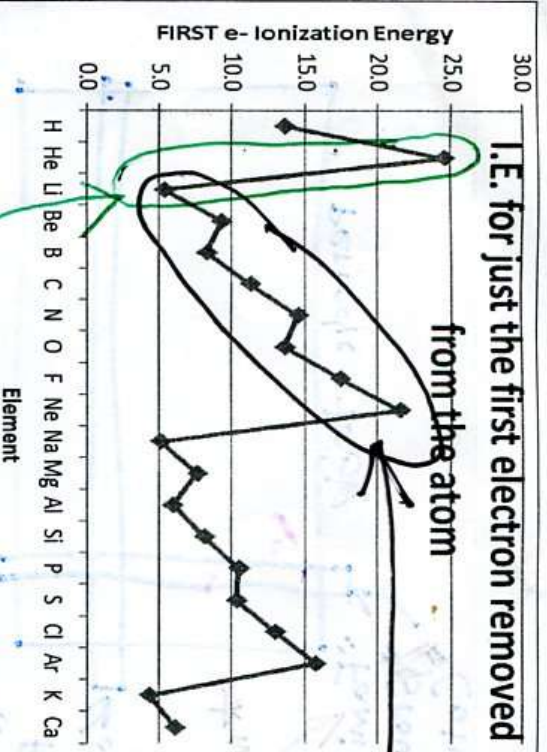


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Ionization Energy for removing all the electrons from Silicon one by one



"Suddenly, ionization energy becomes much higher because ..."

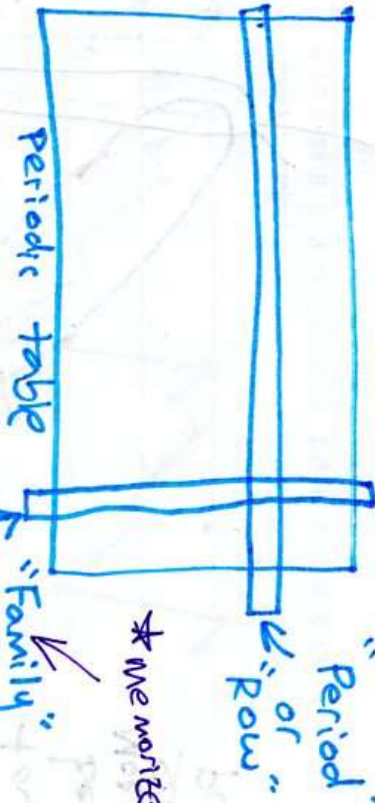
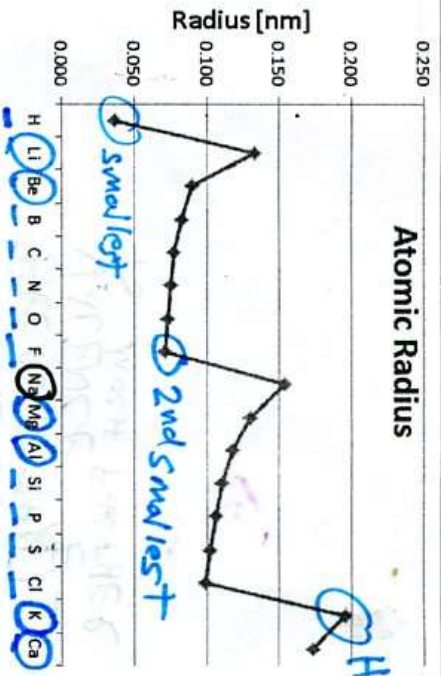


I.E. for just the first electron removed

from the atom

IONIZING gets tougher because more protons, so, more pull

IONIZING GETS EASIER BECAUSE THAT VALENCE IS MUCH FARTHER



* we write

Periodic table

Family Group Column

Ionization Energy for removing all the electrons from SODIUM, one by one

