

# NOTES TO MEMORIZE - DON'T TURN IN

wArMuP - from the indicated pages, fill in each definition below:

- a. (p.398) "The atomic radius is..."
- b. (p. 405) "The electronegativity of an element is..."
- c. (p. 401) "The energy required ..."

**PAID**  
MAY 00 2015

BY: .....

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Three points classwork credit: Get a stamp here

The top half of page 406 has a great cheat sheet to summarize up Periodic Trends. Copy it here and memorize it by next Friday. It will help A LOT on The Camouflage Worksheet tonight too.

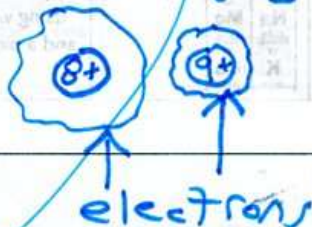
Which element has a smaller radius?

2	He
8	Ne
16	Ar

Fluorine is smaller

Explain why, using words and a picture

Fluorine has more protons so it pulls all the e<sup>-</sup> closer



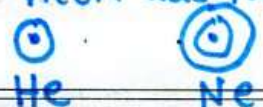
Which element has a smaller radius?

Period 2		Helium
Oxygen	Fluorine	Neon
Sulfur	Chlorine	Argon

Helium is smaller

Explain why, using words and a picture

Helium has one layer of  $e^-$   
But neon has two layers



Which element has a higher first ionization energy?

Period 3		Lithium	Beryllium
Sodium	Magnesium		
Potassium	Calcium		

calcium has a higher I.E.


Explain why, using words and a picture

calcium has more protons.

Which element has a higher first ionization energy?

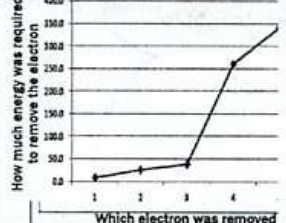
Period 2		Helium
Oxygen	Fluorine	Neon
Sulfur	Chlorine	Argon

Explain why, using words and a picture



Neon Helium

For this element, explain why the first three ionization energies (I.E.'s) are the way they are.



Explain why, using words and a picture

Which element has a greater electronegativity?

Period 2		Helium
Oxygen	Fluorine	Neon
Sulfur	Chlorine	Argon

Explain why, using words and a picture

Which element has a greater electronegativity?

Period 3		Lithium	Beryllium
Sodium	Magnesium		
Potassium	Calcium		

Explain why, using words and a picture