

Make up tests: Come at lunch and after school to take missing tests. No appointment necessary.

Last day for late anything is a week from tomorrow (June 5).

the  
exam  
covers

Second semester only

The Final Exam is the second week of June. Start making your cheat sheet. It must be one sided, hand written, 8½" x 11"

I will post a jumbo review packet online this Sunday at [genest.weebly.com](http://genest.weebly.com). We will be teaching new material every day. Next Week we will also start reviewing.

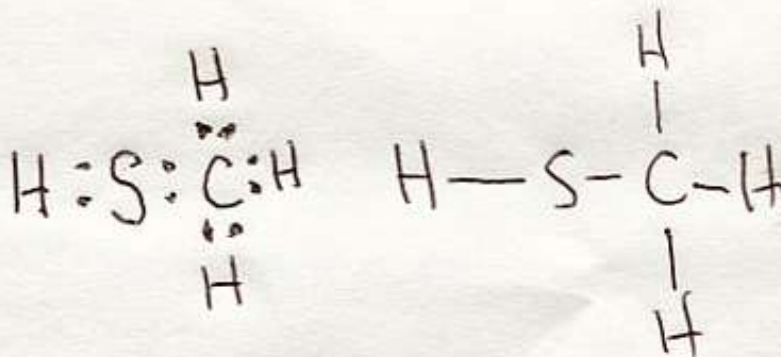
Bring your textbook back! Counts as a ten point homework assignment.

10 pts if you bring back your book this week

7 pts if any day next week.

Purpose: How to predict the stability of covalent bonds.

Warmup: First, copy this drawing. Then convert it into a dots-only drawing.



2 What bonds atoms together?

a. BONDS are the very strong force that holds atoms together. Bonds are too strong to pull apart by hand.

b. METALLIC BONDS form between metal and metal

c. IONIC BONDS form between metal and nonmetal

d. COVALENT BONDS form between nonmetal and nonmetal

e. You try it. Predict the type of bond that will form:

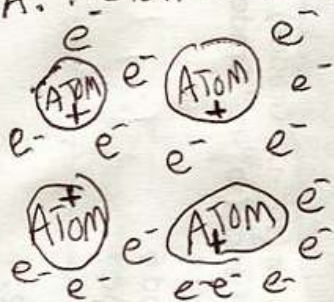
F and Fe form IONIC

2) N and N form covalent

Cu and Cu form metallic

3. WHY DO  $e^-$  bond atoms to each other?

A. metallic Bonds



WITHIN THE SEA  
OF  $e^-$  OPPOSITE  
CHARGED ATOMS ARE  
ATTRACTED

## B. IONIC BONDS

WHEN ATOMS TURN INTO IONS  
WITH OPPOSITE CHARGES:

Before



After

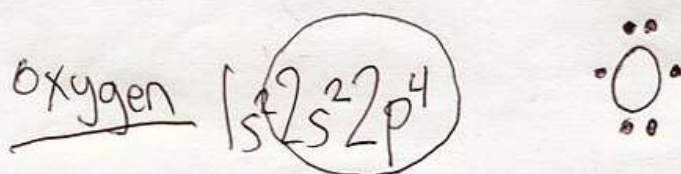
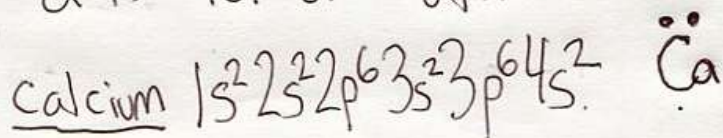


THE OPPOSITES ATTRACT  
AND THIS FORMS A BOND.

## C. COVALENT BONDS

TWO ATOMS SHARE  $e^-$   
and the negative charge  
of the  $e^-$  attracts  
the atoms

4. How to draw the dots for one atom.



5. ATOMS ARE UNSTABLE WHEN THEY DON'T HAVE EIGHT VALENCE  $e^-$

(exception: TWO VALENCE  $e^-$  MAKES SMALL ATOMS STABLE, LIKE Be, Li, He, H)

