

The following people have 0 / 100 on a major test:

Test 5

RC, sd, sl, AJ,

Test 6

CC, na, sd, rh, fj, jm, bs, as, MA, KF, AJ, SK

Test 7

JDV, KE, na, ab, rh, aj, sl, ts, bt, JC, KF, TI, AJ, SK, FP

The last day for fixing late anything is this Thursday (June 5).

The Final Exam

is the second week of June. Start making your cheat sheet. It must be one sided, hand written, 8½" x 11"

The Final Exam covers second semester only.

Look online. genest.weebly.com has a 60 page review packet!

Bring your textbook back! This counts as a ten point homework assignment.

7 / 10 pts if you bring back your book this week

Property	honey	orange Kool-aid
tastes		
color		

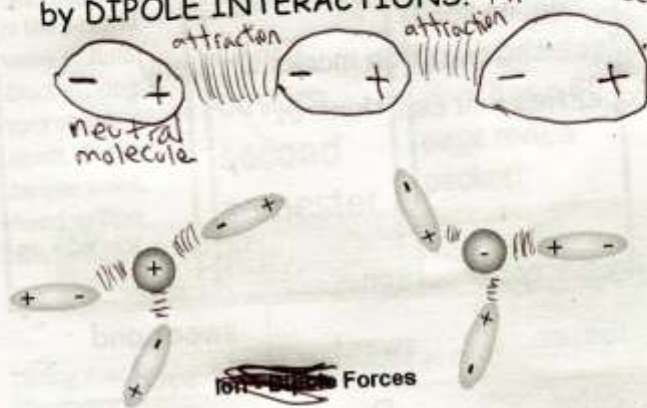


Purpose:

Why is one substance stickier?

1. IMF (intermolecular force) is the technical term for stickiness

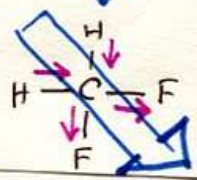
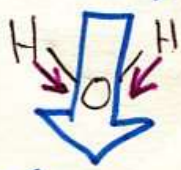
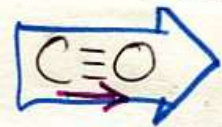
2. Molecule stickiness (IMF) is caused by DIPOLE INTERACTIONS. Three molecules



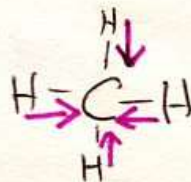
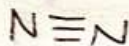
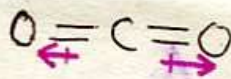
3. Three causes of DIPOLE INTERACTIONS

a. POLAR MOLECULES (sometimes called 'dipoles')

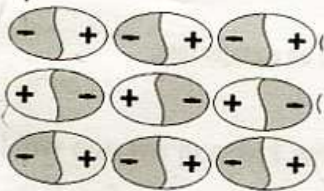
these molecules are polar
because the e^- density of polar bonds add up to an overall polar molecule



these are not polar
Because the polar bonds cancel out



#4



dipole interactions occur because **POSITIVE CHARGED** parts of molecules are attracted to the **NEGATIVE** charged parts of other molecules

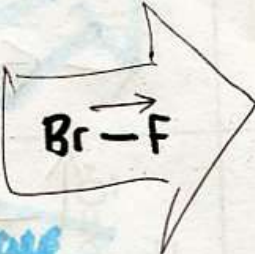
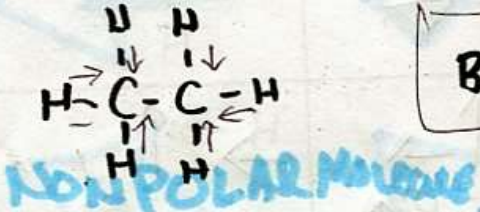
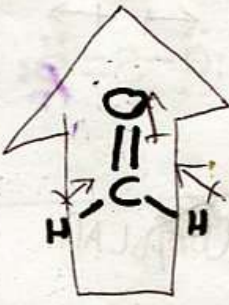
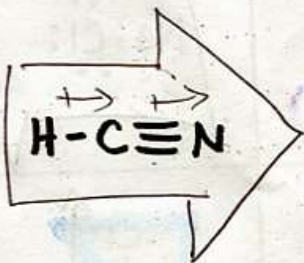
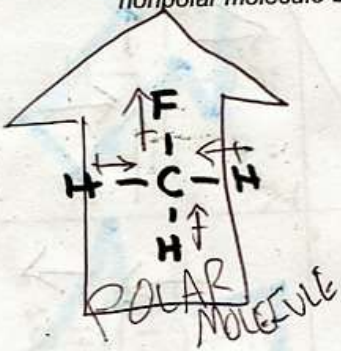
A N S W E R

dipole molecules
 Chemistry: <http://genest.weebly.com>
 Stop in for help every day at lunch and Tues, Wed, & Thurs after school!
 After-hours questions? Email me at home:
ecogover@madison.k12.wi.us



Name _____
 Period _____

- Which element in PERIOD 3 has the least electronegativity? **Na**
- Determine whether each molecule is a dipole by doing the following steps:
 - look up the electronegativity number on your chart from Thursday. Write this number next to each atom in your molecule (this step is optional)
 - draw an arrow to show the direction of polarity of each bond.
 - draw a hollow arrow to show the overall polarity of the molecule OR write *nonpolar molecule* below the molecule



- (Circle one) **Ionic bonds** are usually formed when nonmetals react with (metals/nonmetals)
- (Circle one) **Covalent bonds** are usually formed when nonmetals react with (metals/nonmetals)
- (Circle one) **Metallic bonds** are usually formed when metals react with (metals/nonmetals)
- draw a cloud that is either symmetrical, lopsided, or really lopsided, with the fattest part of the cloud going towards the more electronegative element. Draw an arrow, parallel to the bond, that shows the direction where the e- density is greatest. *The first has been done as an example.*

Continued from front

a. Li-S	b. O-C	c. C-C

- 7 Determine whether each molecule is a dipole by doing the following steps:
- look up the electronegativity number on your chart from Thursday. Write this number next to each atom in your molecule (this step is optional)
 - draw an arrow to show the direction of polarity of each bond.
 - draw a hollow arrow to show the overall polarity of the molecule OR write *nonpolar molecule* below the molecule

<p>NON POLAR molecule</p> $\begin{array}{c} \leftarrow \quad \rightarrow \\ \text{:}\ddot{\text{O}}-\text{S}=\ddot{\text{O}}\text{:} \end{array}$	$\text{Na}:\ddot{\text{Cl}}:$	
<p>Nonpolar</p> $\text{H}-\text{C}\equiv\text{C}-\text{H}$		
<p>NON POLAR molecule</p> $\text{:}\ddot{\text{O}}=\ddot{\text{O}}\text{:}$	<p>POLAR molecule</p>	