Test handed back at the end.

Have you started making your cheat sheet yet?

Purpose: REVIEW OUR FEBRUARY NOTES.

Warmup: Write the formula of each

Gold III Nitrate $Au(NO_3)_3$

Sodium Nitrate $NaNO_3$

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HINTS IN CLASS for Solving window washer

[redacted]

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Go through the eight formulas below and circle any that have a metal with an unpredictable charge Then name each compound. You only use a Roman numeral if there is a metal with unpredictable charge.

1. Na_2CO_3 ____________
2. VCO_3 Vanadium (II) carbonate
3. $Fe(NO_3)_2$ ____________
4. $(NH_4)_2CO_3$ ____________
5. $Au(NO_3)_3$ gold (III) nitrate
6. $Fe_3(PO_4)_2$ ____________

7. What's the formula of each?

Copper (II) bromide ____________

Look at these naming examples to get you warmed up. Notice the asterisks match the footnoted rules		
NaBr is named sodium bromide *[1]	$Sc(OH)_3$ is named scandium hydroxide ** [2]	$V_2(SO_4)_3$ is named vanadium (III) sulfate *** [3]

READ THE ASTERISKS AT THE PAGE BOTTOM. THEY ARE CLUES FOR NAMING THE FOLLOWING COMPOUNDS:

*** $Ti(SO_4)_2$ ____________

*** $FePO_4$ ____________

☆ * NaBr Sodium Bromide

☆ ** $Ca(C_2H_3O_2)_2$ Calcium Acetate

* K_3N ____________

☆ *** CuOH Copper (I) Hydroxide

* $Zn(NO_2)_2$ ____________

V_2S_3 ____________

Ca_3P_2 ____________

[1] *This has one metal and one nonmetal element. name it ELEMENT + ELEMENT + IDE

[2] ** This has three or more elements. You MUST use the polyatomic names from the *back* of your periodic table handout

[3] *** This has a metal element with unpredictable charge, from the middle of the periodic table. You must assign a Roman Numeral. Don't be goofy: Roman Numerals DON'T tell how many atoms, they tell the "plus charge" of a single atom Example, in TiO_2, the name is Titanium(IV) oxide. The "IV" means there is a plus four charge on the metal atom.

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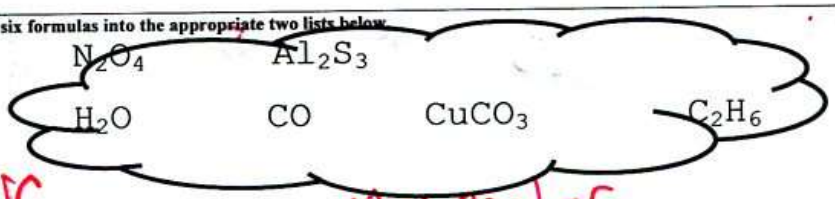
Write the formula of each compound shown below

8. sodium phosphate	11. magnesium nitrite $Mg(NO_3)_2$
9. copper (II) nitrate $Cu(NO_3)_2$	12. tin(IV) oxide
10. copper (II) nitride	13. Aluminum Iodide AlI_3

Instructions: Copy these six formulas into the appropriate two lists below



IONIC Molecular

List each ionic formula then write a name using the rules from last week:	List each molecular formula, then write a name using the mono/di/tri rules we learned today:
1) Al_2S_3 aluminum sulfide	3) CO carbon monoxide
2) $CuCO_3$ copper(II) carbonate	4) H_2O
	5) C_2H_6 6) N_2O_4

Instructions:

A) Circle any substance that is a molecular substance

B) Name each molecular substance you circled USING mono, di. tri RULES.

14. CBr_4
15. Hg_2O
16. NH_3
17. CsBr
18. AgF
19. SnI_2
20. N_2O
21. GeH_4
22. N_2Br_4
23. P_2S_5
24. SeO_2
25. HgS
26. CuI

C) Now go back and name each ionic substance USING LAST WEEK'S RULES