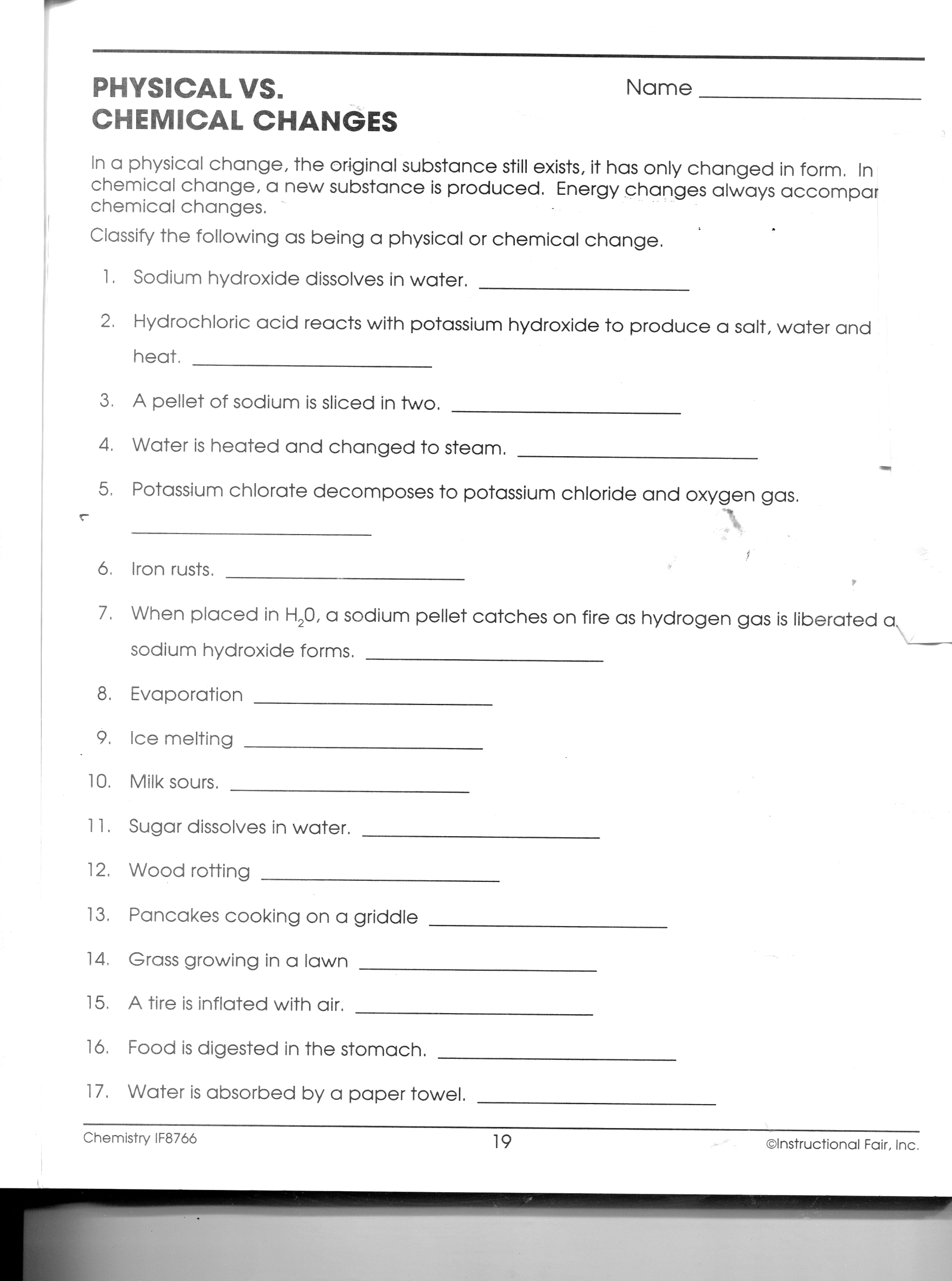
|  |  |  |
| --- | --- | --- |
| ***Physical Changes vs Chemical Changes***  East HS Chemistry  Mr.Genest |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  Period \_\_\_\_\_\_\_\_\_\_ |

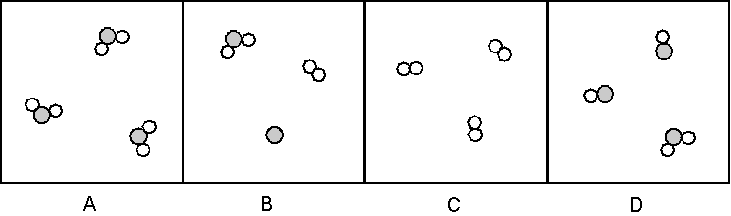


1. These balloons both contain NH3. Inside the balloon on the right draw an appropriate number of NH3.molecules.

|  |  |
| --- | --- |
|  |  |
| 5 liter balloon | 10 liter balloon |
|  |  |

.

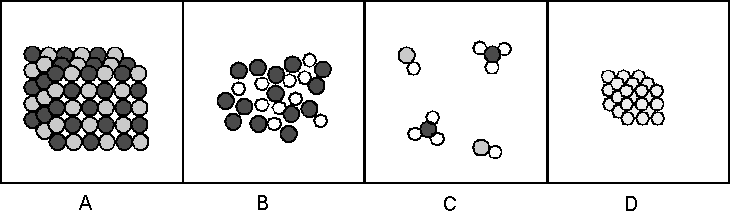
1. Consider the four containers below.



a. Which of these are mixtures? pure substances?

b. Which contain only compounds? only elements

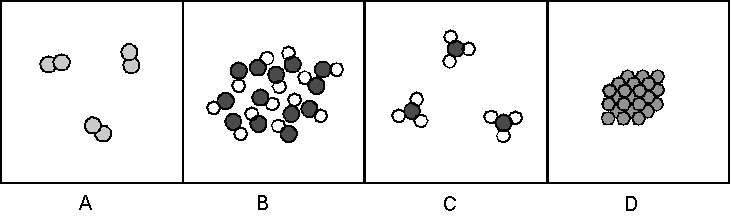
1. Consider the four containers below.



a. Which of these are mixtures? pure substances?

b. Which contain only compounds? only elements

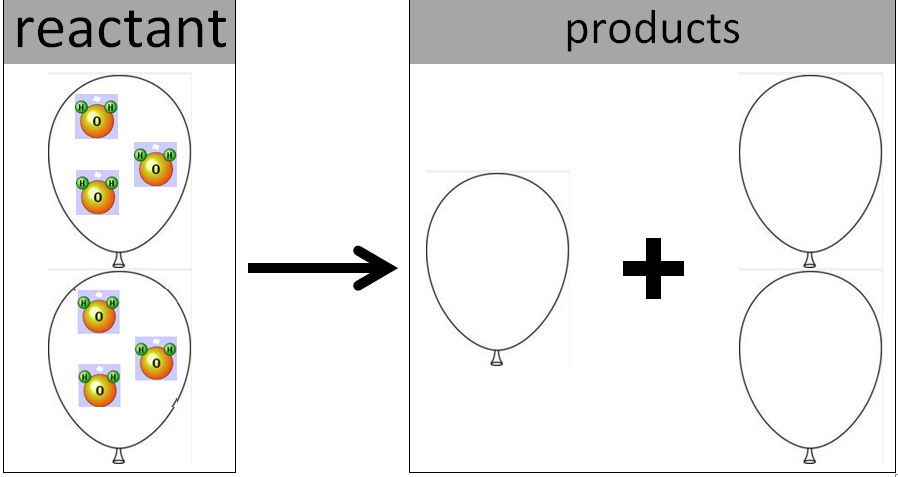
1. Consider the four containers below.



a. Which of these are mixtures? pure substances?

b. Which contain only compounds? only elements

1. Which of the containers in #21 contain a gas? a liquid a solid



1. A chemist took two balloons full of gas. She set them on fire. They created two different products. Product A is a substance that filled one balloon. Product B is a substance that filled two balloons. Draw molecules inside the three balloons showing what she created.

Hints:

* Avogadro’s Principle: Equal sized boxes have the same number of gas molecules.
* The Law of Conservation of Mass: Any atoms that exist before the arrow must be the same as the number of atoms after the arrow.
* Assume all boxes are the same size.