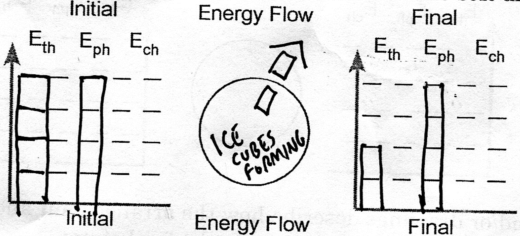
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| Review #2 for the Jan 20th/ Jan 21st Final Exam  East.H.S. ©λ€M|5+rγ |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  Date\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  visit http://genest.weebly.com |

1. If you have 7030. kg, how many Gg do you have? (you should write some converter on your ‘cheat sheet’)
2. If you have 0.005011 kiloliters, how many microliters do you have?
3. We describe three storage “accounts” to understand the changes we see in chemistry. State their names and describe how energy is stored in these three storage modes
4. If a piece of metal that weighs 55.88 grams has a specific heat of 0.660 joules/gram-degreeC and it cools from 877°C to 25°C, how many joules does it lose? Start by writing the formula for Q from your memory or a cheat sheet.
5. A student filled a graduated cylinder with water and read the meniscus at 25.8 mL. The student then dropped a solid material into the graduated cylinder and the water level rose to 35.9 mL. If the solid material had a density of 2.99 g/mL, determine the mass of the solid object.
6. A tray of water (20 ˚C) is placed in the freezer and turns into ice cubes (- 8 ˚C)



List one or more **mistakes** in the solution shown above

1. The numbers below will NOT be given to you on your final exam. You will only be given a periodic table. **Instructions**: Write down some of these and any numbers like these onto your ‘cheat sheet’ now

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| ***You will NOT be given these numbers on the final exam.*** |  |
| 0 degrees C = 273 kelvins do you know the definition of absolute zero?  760. torr = 760. mmHg = 1.00 atm = 101 kPa = 101,300 pascals = 14.7 p.s.i.  1.00 mole of things is 6.02x1023 things.  *\*  ***\*\* you should either memorize your metric units or commit them to your sheet*** | |

1. How many mm Hg are in 77.3 kPa?
2. How many torr are in 1.19 atm?

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| This is a before and after problem. Your solution should be a ‘lonely number’\* followed by one or more ratios. The ratios will have the same units on top and bottom. |
| 1. A quantity of gas exerts a pressure of \*98.6 kPa at a at a temperature of 22.0 °C. If the volume remains unchanged, what pressure will it exert at -8.0°C? |

1. If a beachcomber finds a copper coin that contains 8 x 1021 atoms of copper, what is the volume of the coin? Assume that 255 atoms of copper have a mass of 2.69 x 10-20 grams. Also assume that the density of copper is 8.98 grams per mL.

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| 1. Is this a substance? (yes / no ) |  |

1. Round each to two significant figures

* 550.9 is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
* 951236 is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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| |  |  | | --- | --- | | **Substance** | **Density (g/mL)** | | Aluminum | 2.70 | | Titanium | 4.54 | | Zinc | 7.13 | | Tin | 7.31 | | Iron | 7.87 | | Nickel | 8.90 | | Copper | 8.96 | | Silver | 10.50 | | Lead | 11.35 | | Mercury | 13.55 | | Gold | 19.30 |   you **don’t** need to put numbers like this on your cheat sheet | 1. You made some cubes out of each metal in the table that each measures 2.00 cm on every side.   a. What is the volume of each cube in **cm3**? in **mL**?    V = \_\_\_\_\_\_ cm3 V = \_\_\_\_\_\_ mL |
| . b. Find the mass of these metal cubes: (Show your work below)  lead cube \_\_\_\_\_\_\_\_\_\_\_\_\_\_  nickel cube \_\_\_\_\_\_\_\_\_\_\_\_\_\_  zinc cube \_\_\_\_\_\_\_\_\_\_\_\_\_\_ | |

1. How many significant figures are in each measurement below?

* 75000 miles \_\_\_\_\_\_\_\_\_\_\_\_\_
* 903.0 g \_\_\_\_\_\_\_\_\_\_\_\_\_
* 800.9520 \_\_\_\_\_\_\_\_\_\_\_\_\_
* 0.0005000050 nm \_\_\_\_\_\_\_\_\_\_\_\_

1. Convert the following into scientific notation:
2. 745
3. 0.00852

Convert the following to “normal” numbers (with no exponents anywhere)

1. 7.2 x 10-2

***Mark (H) or (T) for each item below indicate whether it applies to HEAT or TEMPERATURE***

1. \_\_\_\_\_ is measured with a cup of water called a Calorimeter
2. \_\_\_\_\_ is just a measure of the average speed of the particle vibrations
3. \_\_\_\_\_ one common unit for measuring this is joules
4. A cube of aluminum with a density of 2.70 g/mL has a volume of 52.8 mL, what its mass? (Hint, find the numbers in a relationship, find the lonely number, find the unit for the answer. Set up words before numbers.)

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| 1. Using the picture on the right, determine   the pressure of the gas if the atmospheric  pressure is 735 mm Hg.   1. If the CH4 gas became colder the height of liquid would    1. still be 30.0 mm    2. be greater than 30.0 mm    3. be less than 30.0 mm 2. If the atmospheric pressure became less (due to bad weather) the height of liquid would    1. still be 30.0 mm    2. be greater than 30.0 mm   be less than 30.0 mm   1. How many torr is the ambient pressure in this story problem? | http://www.chemteam.info/GasLaw/manometer-for-AP-worksheet.gif |

1. Write your best estimate of the length of this object using

* The left ruler: \_\_\_\_\_\_\_\_\_\_\_\_\_\_
* The right ruler: \_\_\_\_\_\_\_\_\_\_\_\_\_\_

