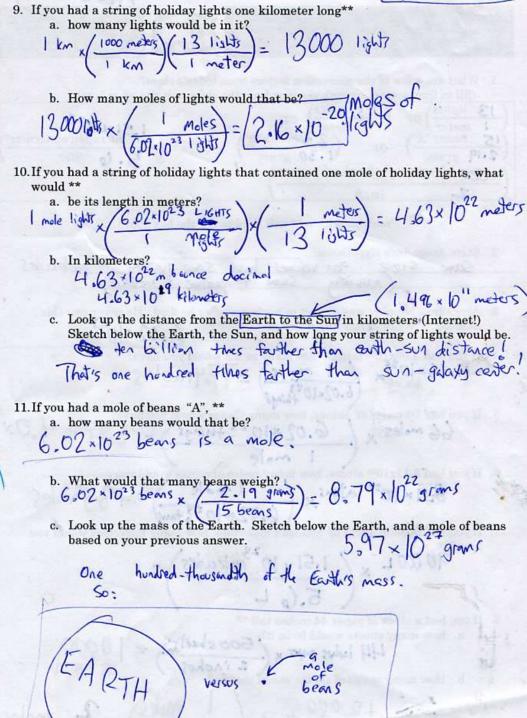
	Avogadro's Number is also called the MOLE! CheMistry: http://genest.weebly.com Stop in for help every day at lunch and Tues. Weds. & Thurs after school! AN SWEIS Name Period
	 What are a few of the conversion factors from today's class? (fill in from your classwork sheet or from the class web site after 5pm)
(-(-	$\frac{13 \text{ lights}}{1 \text{ meter}} \text{ or } \left(\frac{13 \text{ lights}}{1 \text{ meter}}\right) \text{ or } \left(\frac{23 \text{ beans "B"}}{1.30 \text{ gram}}\right) \text{ or } \left(\frac{23 \text{ beans "B"}}{1.30 \text{ gram}}\right) \text{ or } \left(\frac{1.51 \times 10^{23}}{1.30 \text{ gram}}\right)$ $\frac{1.51 \times 10^{23}}{1.30 \text{ gram}}$
1	2 inch
	Problems with a ** will need to use the box above
	2. State Avogadro's Hypothesis: Same SIZE gas volumes have some quantity of molecules (assuming same temperature pressure)
	3. Avogadro's Number was named in his honor after he died. He never saw this number. Write Avogadro's number: 6.02 • 10 ²³
1	4. If you had 4.66x10 ²⁵ things, how many moles would you have? H.66x10 ²⁵ things 1 mole 77.4 moles 6.02×10 ²³ things 1 moles 1 mole
	5. If you had 66 moles of things, how many things would you have? 66 moles x 6.02 × 10 ²³ things = 3.972 4.0 × 1 mole
	6. If you had 8.11x10 ²⁵ atoms, how many moles of atoms would you have? 8.11x10 ²⁵ atoms x (1 no le) = 135 moles
	7. If you had 90,000. liters of nitrogen molecules, how many molecules would you
	1.51 x 10 ²³ molecules =
*	8. If you had a stack of paper 44 inches tall ** a. how many sheets would be in it? HU lacks paper x (2 inches) = 0000
	b. How many moles of sheets would that be?
	10 000 x (1 moles) = 2 moles
	(20013



180