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| Molecular Mass PRacticeCλeMis+ry Final exams start January 20. YOU SHOULD START WRITING YOUR ONE PAGE ‘CHEAT SHEET’ TODAY. See the website.  | http://www.disneyclips.com/imagesnewb/imageslwrakr01/fev32.gif | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| 1. For the substance dipropylene glycol C3H8O2 (an ingredient in Old Spice deodorant)…
 |
| According to the periodic table, what is the mass of a mole of this molecule? \*\* |
| Find the mass of 0.0550 moles of this molecule. |

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| 1. For the substance iodine (Remember the Wacky Seven? Remember Hoffen Brickl?)
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| How would you draw one molecule? | According to the periodic table, what is the mass of a mole of this molecule? \*\* |
| Find the number of molecules in 33.77 grams of this molecule. |

1. How many moles of copper are in 4.7 x 1022 atoms of copper?
2. How many moles of molecules are in each of the following?

 a. 1.50 x 1023 molecules on NH3

 b. 6.02 x 1022 molecules of Br2

1. What is the mass in grams of each of the following?

 a. 0.720 moles of N2O

 b 5.08x1021 molecules of C2H6O

1. How many **hydrogen atoms** are in a molecule of each of these substances?

a. Ca(OH)2 \_\_\_\_\_\_\_ b. C3H8O \_\_\_\_\_\_\_\_ c. (NH4)3PO4 \_\_\_\_\_\_\_ d. HC2H3O \_\_\_\_\_\_\_\_

1. (Challenge problem!)How many atoms of chlorine are there in 16.5 g of iron (III) chloride, FeCl3?