

Name _____

Trade-n-Grade Fake Quiz,

THIS IS A GREAT SUMMARY OF EVERYTHING
YOU NEED TO KNOW TO WRITE A CORRECT
NAME

1. In the substance K_3N , what does the three mean?
2. When naming, always start by chopping the formula into a left and right half. The Left half is almost always a (Metal / Nonmetal). The left half is always a (Cation / Anion)
3. The left and right half of an ionic substance formula have (same / opposite) charges.
4. The charges in any ionic compound must total up to make a substance that is (cationic / neutral / anionic)
5. The left half is seldom polyatomic. According to our table the only two polyatomic cations we will use have the following formulas: _____ and _____

6. To name the left half, write element name and, if necessary, a roman numeral. List two elements that would require a roman numeral if they were on the left half: _____ .
7. The left half needs no roman numeral if it has a predictable charge. List two elements that would not require a roman numeral _____ .

To find the roman numeral you need to figure out what charge is on the metal. Answer the following questions about this

8. Metals are on the (left / right) side of the periodic table.
9. The total charge of any compound must be (+ / zero / -)
10. Based on #9, Choose a charge for the metal in the compound PbO_3 _____
11. To name the right side of a formula, if it is a poly atomic ion , look up the name on your list. Circle which of these are polyatomic ions: $Na_2Cr_2O_7$ NH_4Cl $MgCl_2$

If the right side is not a polyatomic the name is element and ide.
For example, give the complete name of KF _____

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- three atoms of potassium
1. In the substance K_3N , what does the three mean?
 2. When naming, always start by chopping the formula into a left and right half. The Left half is almost always a (Metal / Nonmetal). The left half is always a (Cation / Anion).
 3. The left and right half of an ionic substance formula have (same / opposite) charges.
 4. The charges in any ionic compound must total up to make a substance that is (cationic / neutral / anionic)
 5. The left half is seldom polyatomic. According to our table the only two polyatomic cations we will use have the following formulas: NH_4^+ and H_3O^+

6. To name the left half, write element name and, if necessary, a roman numeral. List two elements that would require a roman numeral if they were on the left half: Anything with unpredictable charge like Fe, Co, Ni, Mo
7. The left half needs no roman numeral if it has a predictable charge. List two elements that would not require a roman numeral The first two columns and the Ag^+ , Zn^{2+} , Al^{3+} area

To find the roman numeral you need to figure out what charge is on the metal. Answer the following questions about this

8. Metals are on the (left / right) side of the periodic table.
9. The total charge of any compound must be (+ / zero / -)
10. Based on #9, Choose a charge for the metal in the compound PbO_3 6^+
11. To name the right side of a formula, if it is a poly atomic ion, look up the name on your list. Circle which of these are polyatomic ions: $Na_2Cr_2O_7$, NH_4Cl , $MgCl_2$

If the right side is not a polyatomic the name is element and ide.

For example, give the complete name of KF potassium fluoride

ANSWERS