|  |  |  |
| --- | --- | --- |
| *Naming Ions.*  CλeMis+ry: http://genest.weebly.com  Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_  Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

1. Look up the following polyatomic ions on the back of your new periodic table. Write down the formula (***including the charge*** ):

ammonium \_\_\_\_\_\_\_\_\_\_ acet**ate** \_\_\_\_\_\_\_\_\_\_ carbon**ate** \_\_\_\_\_\_\_\_\_\_

dichrom**ate** \_\_\_\_\_\_\_\_\_\_ hydroxide \_\_\_\_\_\_\_\_\_\_ nitr**ate** \_\_\_\_\_\_\_\_\_\_

oxal**ate** \_\_\_\_\_\_\_\_\_\_ sulf**ate** \_\_\_\_\_\_\_\_\_\_ phosph**ate** \_\_\_\_\_\_\_\_\_\_

1. What element do most of the polyatomic ions have in the formula? \_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. What type of elements are found in the polyatomic ions? (metal/nonmetal)

|  |  |  |
| --- | --- | --- |
| Look at these naming examples to get you warmed up. Notice the asterisks to the footnoted rules | | |
| NaBr is named **sodium bromide \*[[1]](#footnote-1)** | Sc(OH)3 is named **scandium hydroxide \*\* [[2]](#footnote-2)** | V2(SO4)3 is named **vanadium (III) sulfate \*\*\* [[3]](#footnote-3)** |

\*\*\* Ti(SO4)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\*\*\* FePO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\* NaBr \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\*\* Ca(C2H3O2)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\* K3N \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\*\*\* CuOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\* Zn(NO2)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

V2S3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Ca3P2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. \*This has one metal and one nonmetal element. name it ELEMENT + ELEMENT + IDE [↑](#footnote-ref-1)
2. \*\* This has three or more elements. You MUST use the polyatomic names from the *back* of your periodic table handout [↑](#footnote-ref-2)
3. \*\*\* This has a metal element with unpredictable charge, from the middle of the periodic table. You must assign a Roman Numeral. Don't be goofy: Roman Numerals DON'T tell how many atoms, they tell the "plus charge" of a single atom Example, in TiO2, the name is Titanium(IV) oxide. The "IV" means there is a plus four charge on the metal atom. [↑](#footnote-ref-3)