Some answers to the review packet, problem number 16 from the second page of the review packet Some answers to the review packet, problem number 16 from the second page of the review packet

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c2) 4.0 L
$$\times \frac{1.1 \text{ atm}}{3.4 \text{ atm}} = 1.3 \text{ L}$$

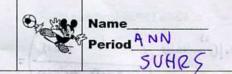
c3

5.01 iters $\times \frac{123 \text{ K}}{473 \text{ K}} \times \frac{1.05 \text{ atm}}{0.65 \text{ atm}} = 2.1 \text{ Lites}$

Molecular Mass PRactice

CaeMis+ry

Final exams start January 20. YOU SHOULD START WRITING YOUR ONE PAGE 'CHEAT SHEET' TODAY. See the website.



For the substance dipropylene glycol C₃H₈O₂ (an ingredient in Old Spice

According to the periodic table, what is the mass of a mole of this molecule? **

C: 3×12.01 = 36.03

H: 8×1.01 = 8.08

O: 2×16.00 = 32.00

76.11 31245

Find the mass of 0.0550 moles of this molecule.

= L1.19 gran 0.0550 mol x

For the substance iodine (Remember the Wacky Seven? Remember Hoffen

How would you draw one molecule?

. 5%

According to the periodic table, what is the mass of a mole of this molecule? **

2x(126.90) = 253,8 grans/mol



Find the number of molecules in 33.77 grams of this molecule.

3. How many moles of copper are in 4.7 x 10²² atoms of copper?

 $\frac{1}{600 \text{ onoles}} = 0.078 \text{ modes}$

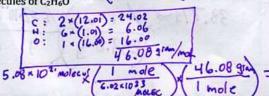
- 4. How many moles of molecules are in each of the following?
 - a. 1.50 x 1023 molecules on NH₃

b.
$$6.02 \times 10^{22}$$
 molecules of Br₂

$$6.02 \times 10^{22}$$
 molecules
$$\times \left(\frac{1 \text{ moles}}{6.02 \times 10^{23}}\right) = 0.100 \text{ moles}$$

5. What is the mass in grams of each of the following?

b 5.08x1021 molecules of C2H6O



6. How many hydrogen atoms are in a molecule of each of these substances?

7. (Challenge problem!)How many atoms of chlorine are there in 16.5 g of iron (III) chloride,