

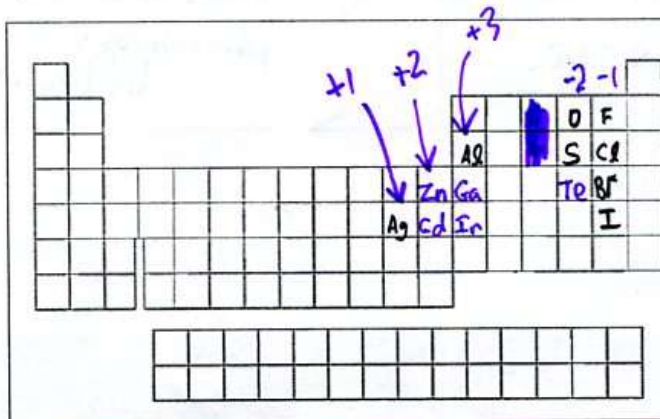
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A N S W E R S



Name _____
 Period _____



- Use your periodic table to write the element symbols and charges for all of the following:

The halogens
 Group 16
 The +1, +2, +3 elements near the middle of the table, that starts with silver

- Fill in just the formula. THEN write the name of the compound.

	zinc	lead (II)	gallium	lead (IV)
chloride	ZnCl ₂ zinc chloride	PbCl ₂ lead(II) chloride	GaCl ₃ gallium chloride	PbCl ₄ lead(IV) chloride
acetate	Zn(C ₂ H ₃ O ₂) ₂ zinc acetate	Pb(C ₂ H ₃ O ₂) ₂ lead(II) acetate	Ga(C ₂ H ₃ O ₂) ₃ gallium acetate	Pb(C ₂ H ₃ O ₂) ₄ lead(IV) acetate
nitrate	Zn(NO ₃) ₂ zinc nitrate	Pb(NO ₃) ₂ lead(II) nitrate	Ga(NO ₃) ₃ gallium nitrate	Pb(NO ₃) ₄ lead(IV) nitrate

- You should only use the Greek prefixes (mono, di, tri, etc) if the compound is (ionic / molecular)
- You should never use the Greek prefixes (mono, di, tri, etc) if the compound (metal / nonmetal)

Name these using the mono, di, tri Greek prefixes

- C₂H₂ dicarbon dihydride
- N₂H₄ dinitrogen tetrahydride
- PH₃ phosphorus trihydride

These are molecular compounds. Name them using the rules for molecular compounds (greek prefixes)

8. N_2O_5 DINITROGEN PENTOXIDE	11. C_3N_2 TRICARBON DINITRIDE
9. P_2O_4 DIPHOSPHOROUS TETROXIDE	12. NI_3 nitrogen triiodide
10. C_2N_3 DICARBON TRINITRIDE	13. PF_2 phosphorous difluoride

read this! Use these asterisks to help you name the compounds in the squares below

¹* This has one metal and one nonmetal element. name it ELEMENT + ELEMENT + IDE


¹** This has three or more elements. You MUST use the polyatomic names from the *back* of your periodic table handout

¹*** This has a metal element with unpredictable charge, from the middle of the periodic table. You must assign a Roman Numeral. Don't be goofy: Roman Numerals DON'T tell how many atoms, they tell the "plus charge" of a single atom. Example, in TiO_2 , the name is Titanium(IV) oxide. The "IV" means there is a plus four charge on the metal atom.

$Mg(ClO_3)_2$ ** magnesium chlorite	NH_4NO_3 ** ammonium nitrate	$NaCl$ * sodium chloride	AgI * silver iodide
$CaSO_4$ ** calcium sulfate	CaO * calcium oxide	PbO *** lead(II) oxide	$Mg(MnO_4)_2$ ** magnesium permanganate
FeS *** iron(II) sulfide	Fe_2S_3 ** iron(III) sulfide	$KMnO_4$ ** potassium permanganate	$Ca(OH)_2$ ** calcium hydroxide
$Mo(ClO_4)_2$ *** molybdenum(II) perchlorate	$CoCl_2$ *** cobalt(II) chloride	$CuBr$ *** copper(I) bromide	$CoCl_3$ *** cobalt(III) chloride

Get a
stamp
when
done:

Draw as circles. Then write the formula of each

1. see	2.	3.
4. the	5. answers	6.
7. below	8.	9. 
10.	11.	12.

When in Doubt, Draw it Out. Answers to this will be online after 5pm

ANSWERS

(1)
Copper (I) sulfate
 Cu_2SO_4

(2)
Copper (II) sulfite
 $CuSO_3$

(3)
dinitrogen tetrafluoride
 N_2F_4

(4)
tricarbon hexahydride
 C_3H_6

(5)
Iron (III) oxide
 Fe_2O_3

(6)
Copper (I) chromate
 Cu_2CrO_4

(7)
Tin (II) sulfate
 $SnSO_4$

(8)
nitrogen monoxide
 NO

(9)
triiodide
 I_3

(10)
Mercury (I) sulfite
 Hg_2SO_3

(11)
Calcium bromide
 $CaBr_2$

(12)
Aluminum oxide
 Al_2O_3

