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| --- | --- | --- |
| *Review #2*CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

# Section 1

1. In our lab experiment, we found two properties that are different for ionic substances versus molecular substances. What were those two properties? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Based on Thomson’s Plum Pudding model, what has to happen to an atom for it to become positive. Draw the Plum Pudding model as part of your explanation.
3. The substance C6H12O6 is ( a molecular substance / an ionic substance )
4. The substance Br2 is ( a molecular substance / an ionic substance )
5. In the formula SO42-
	1. What does the 2 mean? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. What does the 4 mean? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. If dropped into water and dissolved, zinc chloride would probably ( conduct electricity / not conduct electricity )
7. In the following pair of substances, circle the one that would be expected to melt at a LOWER temperature:
8. Br2 or NaBr
9. In the following pair of substances, circle the one that would be expected to melt at a LOWER temperature:

C6H12O6 or KNO3

# Write the name of each of the following compounds.

# CF4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# NO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# N2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# SO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

#

# Write the chemical formula for each of the given names.

# nitrogen triiodide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# dinitrogen tetrabromide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# diboron hexahydride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# carbon monoxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

# go online. the class web page has plenty more practice problems!

# Section 2

Wow, the answers are here. How convenient! Cover up the answers and peek at them when necessary.

|  |  |  |  |
| --- | --- | --- | --- |
| Fe2(Cr2O7)3 | FeCrO4 | Fe2(CrO4)3 | CoP |
| Na2SO4 | NiF2 | Be(NO2)2 | Be(NO3)2 |
| Mo(OH)3 | K2C2O4 | Hg2Cl2 | HgCl2 |



Write the formula

|  |  |  |  |
| --- | --- | --- | --- |
| Tin (IV) carbonate | Tin (II) bicarbonate | Calcium phosphate | Iron (III) hydroxide |
| Sodium hydroxide | Potassium hydroxide | Iron (II) hydroxide | Titanium (III) sulfate |
| Manganese (IV) oxide | Calcium fluoride | Iron (III) sulfite | Iron (II) sulfite |

