

## Review #2

Cl@Mis+ry: <http://genest.weebly.com>

Stop in for help every day at lunch and Tues & Thurs after school!



Name \_\_\_\_\_

Period \_\_\_\_\_

The test can ask anything from notes or homework since January 26 so memorize those from your notebook and at the class website listed above. See the list of topics on the back of this sheet. Naming is only about half of the test. To practice that half, here are some practice problems...with answers:

$\text{Fe}_2(\text{Cr}_2\text{O}_7)_3$	$\text{FeCrO}_4$	$\text{Fe}_2(\text{CrO}_4)_3$	$\text{CoP}$
$\text{Na}_2\text{SO}_4$	$\text{NiF}_2$	$\text{Be}(\text{NO}_2)_2$	$\text{Be}(\text{NO}_3)_2$
$\text{Mo}(\text{OH})_3$	$\text{K}_2\text{C}_2\text{O}_4$	$\text{Hg}_2\text{Cl}_2$	$\text{HgCl}_2$
$\text{Zn}(\text{H}_2\text{PO}_4)_2$	$\text{Fe}_2\text{S}_3$	$\text{Zn}(\text{HS})_2$	$\text{Ag}_2\text{SO}_3$
$\text{Pb}_3(\text{PO}_4)_4$	$\text{Zr}(\text{OH})_4$	$\text{Ag}_2\text{S}$	$\text{Ag}_2\text{SO}_4$
$\text{K}_3\text{N}$	$\text{ZrPO}_4$	$\text{Li}_2\text{O}$	$\text{Ag}_2\text{S}_2\text{O}_3$

$\text{Fe}_2(\text{Cr}_2\text{O}_7)_3$ iron(III) dichromate	$\text{FeCrO}_4$ iron(II) chromate	$\text{Fe}_2(\text{CrO}_4)_3$ Iron(III) chromate	$\text{CoP}$ cobalt(II) phosphide
$\text{Na}_2\text{SO}_4$ sodium sulfate	$\text{NiF}_2$ nickel(II) fluoride	$\text{Be}(\text{NO}_2)_2$ beryllium nitrite	$\text{Be}(\text{NO}_3)_2$ beryllium nitrate
$\text{Mo}(\text{OH})_3$ molybdenum(III) hydroxide	$\text{K}_2\text{C}_2\text{O}_4$ potassium oxalate	$\text{Hg}_2\text{Cl}_2$ mercury(I) chloride	$\text{HgCl}_2$ mercury(II) chloride
$\text{Zn}(\text{H}_2\text{PO}_4)_2$ zinc dihydrogen phosphate	$\text{Fe}_2\text{S}_3$ iron(III) sulfide	$\text{Zn}(\text{HS})_2$ zinc hydrogen sulfide	$\text{Ag}_2\text{SO}_3$ silver sulfite
$\text{Pb}_3(\text{PO}_4)_4$ lead(IV) phosphate	$\text{Zr}(\text{OH})_4$ zirconium(IV) hydroxide	$\text{Ag}_2\text{S}$ silver sulfide	$\text{Ag}_2\text{SO}_4$ silver sulfate
$\text{K}_3\text{N}$ potassium nitride	$\text{ZrPO}_4$ zirconium(IV) phosphate	$\text{Li}_2\text{O}$ lithium oxide	$\text{Ag}_2\text{S}_2\text{O}_3$ silver thiosulfate

I one  
 II two  
 III three  
 IV four  
 V five  
 VI six

Tin (IV) carbonate	Tin (II) bicarbonate	Calcium phosphate	Iron (III) hydroxide
Sodium hydroxide	Potassium hydroxide	Iron (II) hydroxide	Titanium (III) sulfate
Manganese (IV) oxide	Calcium fluoride	Iron (III) sulfite	Iron (II) sulfite
Ammonium nitrate	Chromium (III) oxide	Chromium (II) acetate	Silver phosphate
Molybdenum (VI) chlorite	Potassium permanganate	Nickel (II) formate	Lead (IV) oxalate
Zinc carbonate	Lead (II) thiocyanate	Cobalt (III) chromate	Mercury (I) dichromate

Tin (IV) carbonate $\text{Sn}(\text{CO}_3)_2$	Tin (II) bicarbonate $\text{Sn}(\text{HCO}_3)_2$	Calcium phosphate $\text{Ca}_3(\text{PO}_4)_2$	Iron (III) hydroxide $\text{Fe}(\text{OH})_3$
Sodium hydroxide $\text{NaOH}$	Potassium hydroxide $\text{KOH}$	Iron (II) hydroxide $\text{Fe}(\text{OH})_2$	Titanium (III) sulfate $\text{Ti}_2(\text{SO}_4)_3$
Manganese (IV) oxide $\text{MnO}_2$	Calcium fluoride $\text{CaF}_2$	Iron (III) sulfite $\text{Fe}_2(\text{SO}_3)_3$	Iron (II) sulfite $\text{Fe}(\text{SO}_3)$ <i>parentheses not required</i>
Ammonium nitrate $\text{NH}_4\text{NO}_3$	Chromium (III) oxide $\text{Cr}_2\text{O}_3$	Chromium (II) acetate $\text{Cr}(\text{CH}_3\text{COO})_2$	Silver phosphate $\text{Ag}_3\text{PO}_4$
Molybdenum (VI) chlorite $\text{Mo}(\text{ClO}_2)_6$	Potassium permanganate $\text{KMnO}_4$	Nickel (II) formate $\text{Ni}(\text{HCOO})_2$	Lead (IV) oxalate $\text{Pb}(\text{C}_2\text{O}_4)_2$
Zinc carbonate $\text{ZnCO}_3$	Lead (II) thiocyanate $\text{Pb}(\text{SCN})_2$	Cobalt (III) chromate $\text{Co}_2(\text{CrO}_4)_3$	Mercury (I) dichromate $\text{Hg}_2\text{Cr}_2\text{O}_7$

## Instructional goals

1. Describe the evidence that supports the idea that the simple particles have a property we call charge.
2. Describe the evidence that led Thomson to suggest that the mobile charge in atoms is negative.
3. Use the Thomson model of the atom to account for the fact that neutral atoms can become either positively or negatively charged by the loss or gain of electrons.
4. List properties that distinguish metals from non-metals.
5. Describe the evidence that distinguishes ionic from molecular or atomic solids.
6. Given the formula of an ionic substance, state its name.
7. Given the name of ionic substance, write its formula.
8. From the name or formula of a substance determine whether that substance is ionic or molecular.