

New notebook system:

SCREEN IS NOTES goes onto right pages

CHALKBOARD IS PRACTICE goes onto left pages

Final Exam scantrons handed back at end.

Come to check the questions either at lunch or

Tuesday/Thursday after school

Day 2, Unit 6, January 26, 2016

Purpose:

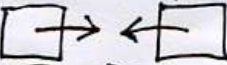


Purpose: Describe the evidence that led Thomson to suggest that the mobile charge in atoms is negative.

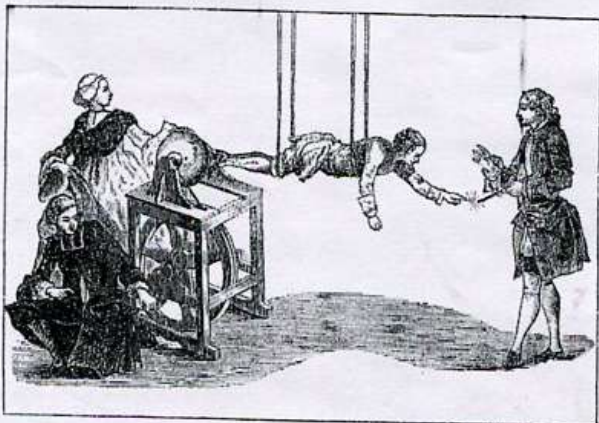
WARMUP :

Fill in the blanks in your notes using yesterday's data table.

"When top tape AND paper are brought close together, they (attract / repel / do nothing)"

1. Vector arrows show the direction of push and pull. Examples:

- attraction 
- repulsion 
- no interaction 



Air	P O S I T I V E
Skin (dry)	
Glass	
Human Hair	
Mica	
Nylon	
Wool	
Cat Fur	
Lead	
Silk	
Aluminum	
Paper	
Cotton	
Steel	
Wood	
Lucite	
Amber	
Rubber Balloon	
Hard Rubber	
Mylar®	
Epoxy glass	
Nickel	
Copper	
Silver	
Gold, Platinum	
Polyester	
Polystyrene	
Orlon, Acrylic	
Polyester	
Cellophane Tape	
Polyurethane	
Polyethylene	
Polypropylene	
Polyimide (Kapton®)	
PVC, Vinyl	
Teflon	
Silicone Rubber	
	N E G A T I V E (-)



2. Rule

Attraction is based on charge.



If two things attract, they have

2. Rule

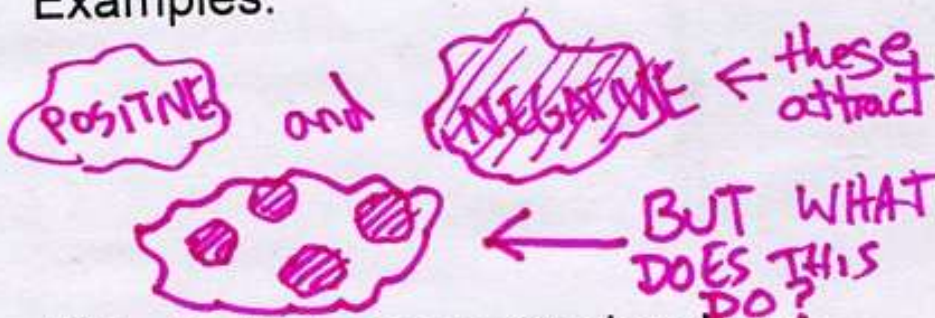
Attraction is based on charge.

If two things attract, they have opposite charge.

If they repel, they have the same charge.

Memorize: "opposites attract"

Examples:



**Neutral things are made when positive and negative are mixed together: IF well mixed, they have no attraction or repulsion:

3. JOHN DALTON imagined Atoms like this



3. What JOHN DALTON thought elements looked like:



4. What JJ THOMSON thought elements looked like:

element X	element Y	element Z
neutral	neutral	neutral
positive!	negative	still neutral

Objects become negatively charged if they gain extra electrons.

Objects become positively charged if they have less than their normal # of electrons.

Home work

Names for Friday's Quiz

East.H.S. ©AEM|s+ry



Name _____

Date _____

visit <http://genest.weebly.com>

The list of which element names to know by Friday is at <http://genest.weebly.com>
It is the same list of elements that you memorized in September.

1. In each case, either supply the missing symbol or the missing name. (memorize these by Friday)

I is iodine

Sn is tin

Pb is lead

Ag is silver

Au is gold

Hg is mercury

Sr is strontium

Ba is barium

U is uranium

Pu is plutonium

Si is silicon

P is phosphorus

2. Write the name of each compound. Use Rule 1 from class today: except in special cases, a name will be "element + element + ide"
example Li_3P is "lithium phosphide"

Ca_3N_2 is calcium nitride

Na_3P is sodium phosphide

KBr is potassium bromide

LiCl is lithium chloride

Na_2S is sodium sulfide

NCl is nitrogen chloride

3. In each case, either supply the missing symbol or the missing name. (memorize

these by Friday)

S is sulfur

Cr is chromium

Cl is chlorine

Mn is Manganese

Ar is argon

Fe is iron

K is potassium

Ni is Nickel

Ca is calcium

Cu is copper

4. Write the name of each compound. Use Rule 1 from class today: except in special cases, a name will be "element + element + ide"

example Li_3P is "lithium phosphide"

CaCl_2 is calcium chloride

HCl is hydrogen chloride

MgI_2 is magnesium iodide

NaBr is sodium bromide

K_3P is potassium phosphide

K_2O is potassium oxide