|  |  |  |
| --- | --- | --- |
| *Naming Ions.*CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

**Name each compound in the next six questions.**

**Situation 1: Metal from first two columns of the table WITH a nonmetal**

**The name will be *Element Name + Element Name + “-ide”***

1. Be3N2
2. Ca3P2

**-Situation 2: Metal from first two columns of the table WITH several nonmetals**

**The name will be *Element Name + cheat sheet name***

1. CaCO3
2. Al(NO3)3

**Situation 3: The metal on the left has a mysterious (variable) charge.**

**The name will be *Element Name + roman numeral that only tells the charge + element name + ide***

1. Hg2S
2. Cr2(CO3)3

 ***The next two questions are about this JJ Thomson apparatus shown below:***



1. What did JJ Thomson conclude was shooting in a line from left to right in this drawing?
	1. positive electrons
	2. negative electrons
2. In this cathode ray tube, line is bending up. Based on the direction the electrons are bending, Which to you think is correct,
3. D is positive and E is negative
4. D is negative and E is positive
5. Both D and E are negative
6. Both D and E are positive.

**Finish the name by writing in the appropriate Roman Numeral**

|  |  |
| --- | --- |
| 1. CrO3 Chromium (\_\_\_\_\_\_) oxide
2. CrCO3 Chromium (\_\_\_\_\_) carbonate
 | 1. Pt(NO3)4 Platinum (\_\_\_\_\_) nitrate
2. PdO2 Palladium (\_\_\_\_\_) oxide
 |

**Next to each write either Ionic Compound or Molecular Compound**

1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ CO2
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ NaOH
3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ CaBr2
4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ H2CO3

|  |  |  |  |
| --- | --- | --- | --- |
| Formula | Metal present? | Which of the four situations above are applicable in this compound? One or more may apply. Mark X for any that apply. | Name the compound: |
| Situation1? | Situation2? | Situation3? |  |
| 1. MnF5
 | (yes/no) |  |  |  |  |  |
| 1. V2(CO3)3
 | (yes/no) |  |  |  |  |  |
| 1. MgS
 | (yes/no) |  |  |  |  |  |
| 1. UF3
 | (yes/no) |  |  |  |  |  |

**For the pairs of elements below, write the formula of the compound that would form when those two elements combine and write the name of the compound.**

1. **Sodium and sulfur**
2. **Magnesium and fluorine**
3. **Beryllium and oxygen**

|  |  |  |  |
| --- | --- | --- | --- |
| Formula | Metal present? | Which of the four situations above are applicable in this compound? One or more may apply. Mark X for any that apply. | Name the compound: |
| 1. Au(NO3)3
 | (yes/no) | 1 | 2 | 3 |  |  |
| 1. Na2O
 | (yes/no) |  |  |  |  |  |
| 1. Fe2(CO3)3
 | (yes/no) |  |  |  |  |  |

**Finish the name by writing in the appropriate Roman Numeral**

1. **CrO3 Chromium (\_\_\_\_\_\_) oxide**
2. **CrCO3 Chromium (\_\_\_\_\_) carbonate**
3. **Pt(NO3)4 Platinum (\_\_\_\_\_) nitrate**
4. **Pd3N2 Palladium (\_\_\_\_\_) nitride**