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| *Five reactions & Review Naming*  CλeMis+ry: http://genest.weebly.com  Stop in for help every day at lunch and Tues &Thurs after school! | **by Edward Carter** | Name\_\_\_\_\_\_\_\_\_  Period\_\_\_\_\_\_\_\_ |

*Predict the products for the following reactions then balance the equation,*

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|  |  | Write the complete reaction as a **balanced** reaction. |
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| (combustion)  C3H6 + O2 🡪 ? + ? |  |  |
| (combination) pure calcium reacting with pure fluorine to just a single product |  |  |
| (combustion)  O2 + C5H12O2 🡪 ? + ? |  |  |
| (single replacement)  Na + CaSO4 🡪 ? + ? |  |  |

**Questions about the lab you finished:**

1. List four possible observable changes that are evidence that a chemical reaction has taken place:
2. How did the flaming splint behave when it was inserted into the tube with CO2 (g)

In what way was this different from the reaction of the H2(g) to the flaming splint?

1. In the reaction of magnesium with oxygen gas, a considerable amount of energy was released. This is an example of an *exothermic* reaction.
   1. What other examples of exothermic reactions did you observe?
   2. Re-write the balanced equation for the reaction of Mg and O2, this time with the term “+ energy” on the appropriate side of the equation. (Hint: exothermic means “heat” “out”)
2. You had to heat the NaHCO3 strongly in order for it to decompose. This is an example of an *endothermic* reaction.
   1. What does this tell you about the energy stored in the reactants compared to the energy stored in the product?
   2. Write the balanced equation for the decomposition of NaHCO3, this time with the term “+ energy” on the appropriate side of the equation. (Hint: endothermic means “heat” “in”)

**Instructions:**

**A) Circle any substance that is a molecular substance**

**B) Name each molecular substance you circled USING TODAY’S RULES.**

1. AgF
2. SnI2
3. Na2O
4. K2S
5. ICl
6. CaBr2
7. BaI2
8. Al2S3
9. N2O