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| *Five Reaction Types (Junior)*CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! | Madam C.J Walker | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

1. Write a balanced reaction for the COMBUSTION of C4H10. If you forgot what four things are in all combustion reactions look at last week's notes!
2. Classify each of the reactions below as one of the following reaction types. Then balance them.

COMBUSTION,

DECOMPOSITION,

COMBINATION,

SINGLE REPLACEMENT,

DOUBLE REPLACEMENT

|  |  |
| --- | --- |
| Type of reaction? | Balance the reaciton |
|  | \_\_\_C2H2(g) + \_\_\_O2(g)  \_\_\_H2O(g) + \_\_\_Co2(g) |
|  | \_\_\_Fe(s) + \_\_\_O2(g)  \_\_\_ Fe2O3 (s) |
|  | \_\_\_Fe + \_\_\_O2 arrow\_\_\_FeO |
|  | \_\_\_NaCl arrow\_\_\_Na + \_\_\_Cl2 |
|  | \_\_\_N2 + \_\_\_H2 arrow\_\_\_NH3 |
|  | \_\_\_HCl + \_\_\_FeS arrow\_\_\_FeCl2 + \_\_\_H2S |

1. Write a balanced reaction for the COMBUSTION of C2H4.

Don’t skip Problem #3!

1. Classify each of the reactions below as one of the following reaction types. Then balance them.

COMBUSTION,

DECOMPOSITION,

COMBINATION,

SINGLE REPLACEMENT,

DOUBLE REPLACEMENT

|  |  |
| --- | --- |
| Type of reaction? | Balance the reaciton |
|  | \_\_\_MgCl2 + \_\_\_AgNO3 arrow \_\_\_Mg(NO3)2 + AgCl |
|  | \_\_\_Fe + \_\_\_CuSO4 arrow\_\_\_FeSO4 + \_\_\_Cu  |
|  | \_\_\_C7H16(g) + \_\_\_O2(g)  \_\_\_H2O(g) + \_\_\_Co2(g) |
|  | \_\_\_P4O10 + \_\_\_H2O arrow\_\_\_H3PO4 |
|  | \_\_\_Fe + \_\_\_H2O arrowFe3O4 + \_\_\_H2 |
|  |  \_\_\_H3PO4 arrow \_\_\_H4P2O7 + \_\_\_H2O  |
|  | \_\_\_P + \_\_\_Cl2 arrow\_\_\_PCl3 |
|  |  \_\_\_Al2(SO4)3 + \_\_\_Ca(OH)2 arrow\_\_\_Al(OH)3 + \_\_\_CaSO4  |