

NO NOTEBOOK.

NOW: WALK AROUND  
THE BACK, COLLECT  
NINE STRIPS.

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AT THE BUZZER:

- Be in seats for
- 1) WARMUP ANSWERS
  - 2) HOMEWORK CHECK

BY THE BELL

- 1) HAND IN YOUR LAB?

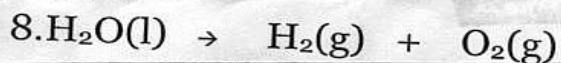
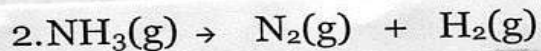
Answers to the warmup activity, which by the way looks similar to tomorrow's Quiz

Partner name:

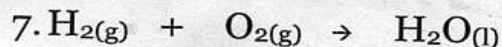
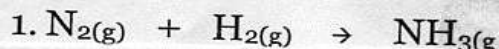
Partner name:

Glue **Tape**, or *Staple* **9** REACTIONS to here

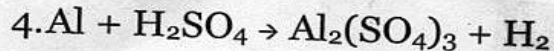
DECOMPOSITION REACTIONS: "One compound falls apart, making two or more new compounds."



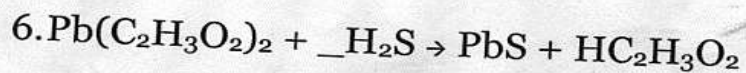
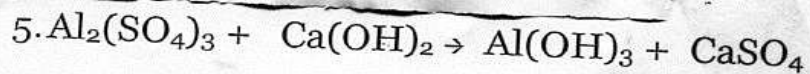
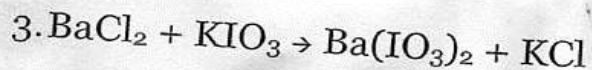
SYNTHESIS REACTIONS: "Two compounds hit and stick, forming one new compound"

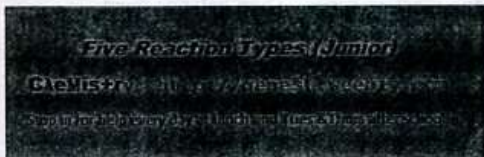


SINGLE REPLACEMENT REACTIONS: "A lonely element cuts in on a compound, and replaces one of the substances after)"



DOUBLE REPLACEMENT REACTIONS: "Two dancing couples trade partners" (Looks like two substances before AND two substances after)





Name \_\_\_\_\_  
Period \_\_\_\_\_

1. Write a balanced reaction for the COMBUSTION of  $C_4H_{10}$ . If you forgot what four things are in all combustion reactions look at last week's notes!



2. Classify each of the reactions below as one of the following reaction types. Then balance them.  
COMBUSTION,  
SINGLE REPLACEMENT,  
DECOMPOSITION,  
DOUBLE REPLACEMENT,  
COMBINATION,

Type of reaction?	Balance the reaction
COMBUSTION	$2 C_2H_2(g) + 5 O_2(g) \rightarrow 2 H_2O(g) + 4 CO_2(g)$
COMBINATION	$4 Fe(s) + 3 O_2(g) \rightarrow 2 Fe_2O_3(s)$
COMBINATION	$2 Fe + 1 O_2 \rightarrow 2 FeO$
DECOMPOSITION	$2 NaCl \rightarrow 2 Na + 1 Cl_2$
COMBINATION	$1 N_2 + 3 H_2 \rightarrow 2 NH_3$
DOUBLE REPLACEMENT	$2 HCl + 1 FeS \rightarrow 1 FeCl_2 + 1 H_2S$

3. Write a balanced reaction for the COMBUSTION of  $C_2H_4$ .



Don't skip Problem #3!



4. Classify each of the reactions below as one of the following reaction types. Then balance them.

COMBUSTION,  
DECOMPOSITION,  
COMBINATION,  
SINGLE REPLACEMENT,  
DOUBLE REPLACEMENT

Type of reaction?	Balance the reaction
DOUBLE REPLACEMENT	$1 \text{ MgCl}_2 + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 2 \text{ AgCl}$
SINGLE REPLACEMENT	$1 \text{ Fe} + 1 \text{ CuSO}_4 \rightarrow 1 \text{ FeSO}_4 + 1 \text{ Cu}$
COMBUSTION	$1 \text{ C}_7\text{H}_{16}(\text{g}) + 11 \text{ O}_2(\text{g}) \rightarrow 8 \text{ H}_2\text{O}(\text{g}) + 7 \text{ CO}_2(\text{g})$
COMBINATION	$1 \text{ P}_4\text{O}_{10} + 6 \text{ H}_2\text{O} \rightarrow 4 \text{ H}_3\text{PO}_4$
SINGLE REPLACEMENT	$3 \text{ Fe} + 4 \text{ H}_2\text{O} \rightarrow \overset{1}{\text{Fe}_3\text{O}_4} + 4 \text{ H}_2$
DECOMPOSITION	$2 \text{ H}_3\text{PO}_4 \rightarrow 1 \text{ H}_4\text{P}_2\text{O}_7 + 1 \text{ H}_2\text{O}$
COMBINATION	$2 \text{ P} + 3 \text{ Cl}_2 \rightarrow 2 \text{ PCl}_3$
DOUBLE REPLACEMENT	$1 \text{ Al}_2(\text{SO}_4)_3 + 3 \text{ Ca(OH)}_2 \rightarrow 2 \text{ Al(OH)}_3 + 3 \text{ CaSO}_4$