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| *Using symbols to show chemical change.*CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| After the formula of each reactant and product you give the **state of the substance in brackets** using the shortform indicated here:* solid **(s)**
* liquid **(l)**
* gas **(g)**
* aqueous **(aq)** This means aqueous which means that the substance was dissolved in water or that it is in solution form.
* A precipitate is an insoluble solid formed during a chemical reaction with solutions. So **(s)** is used.
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1. **When heated, solid carbon reacts with oxygen gas (wacky seven!) to form carbon monoxide (C1O1 or simply CO) gas.**
2. **Aqueous ammonium chloride reacts with aqueous sodium hydroxide to form ammonia gas (NH3), liquid water, and aqueous sodium chloride.**
3. **Aqueous sodium phosphate and liquid water are formed when aqueous sodium hydroxide reacts with phosphoric acid(H3PO4).**
4. **Solid carbon (see #1) reacts with oxygen gas to form carbon dioxide gas.**
5. **Solid sodium reacts with liquid bromine to form solid sodium bromide.**
6. **When heated, solid potassium chlorate yields solid potassium chloride and oxygen gas (see #1).**