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| *Writing Unbalanced Reaction Equations* CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! | **Miles Davis** | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

**Write the unbalanced equations for the following chemical reactions.**

**Write formulas (like H2O) *and* phases (like s, l, g, aq):**

1. When dissolved barium chloride reacts with dissolved potassium sulfate in water, barium sulfate precipitate and aqueous potassium chloride are made.
2. When calcium chloride and potassium phosphate are dissolved in water they react to form aqueous potassium chloride and calcium phosphate powder.
3. When sucrose (C12H22O11) burns in oxygen, carbon dioxide, water and heat are produced.
4. Does the Law of Conservation of Mass agree with what you wrote in today's warmup? Explain.

1. The mass of the candle decreased yesterday. What happened to the wax?
2. Brainstorm two lists: chemicals destroyed by a burning candle and chemicals created by a burning candle. (You should come up with four to ten chemicals, total).

1. Come up with a particle picture that shows the chemicals in a burning candle flame. You must have four or more substances.

1. With the help of question #2 from yesterday, write an unbalanced chemical reaction for the burning of a candle.

**Write the unbalanced equations for the following chemical reactions.**

**Write formulas (like H2O)*and* phases (like s, l, g, aq):**

1. When dissolved calcium hydroxide reacts with sulfuric acid (H2SO4), a precipitate of calcium sulfate, water, and heat are formed.
2. When sodium metal reacts with iron (III) chloride, iron metal and sodium chloride are formed.

**Go back and double check to make sure you put phase letters (s, l, g, aq) on your answers to 1, 2, 3, 9, & 10**

**Also, we e need famous faces to put on top of our worksheets. Please go to Lab Data Blog at http:genest.weebly.com to submit a famous figure from Black History. I’ll give extra credit if we use your choice!**