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| *Writing Names of Molecular Substances*  CλeMis+ry: http://genest.weebly.com  Stop in for help every day at lunch and Tues &Thurs after school! | return.png | Name\_\_\_\_\_\_\_\_\_\_\_\_\_  Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| A.  Definition: Ionic Substances contain a cation and anion. They usually have a metal with one or more nonmetals. Rare exception to this:  Occasionally they have NH4+ instead of a metal. (For example, NH4NO3 is an ionic compound.)  Examples of ionic substances:  \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_ |  | B.  Definition: Molecular Substances contain only nonmetals.  Examples of molecular substances:  \_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_ |
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| C.  How to name Ionic Substances:  Use all the rules we learned on our last test. |  | D.  How to name Molecular Substances:    *mono/di/tri element*  *mono/di/tri element*  *ide*  For example,  N2H5 is named dinitrogen pentoxide  and BrF is named bromine monofluoride  Don’t use Roman numerals!  Omit “mono” if it is at the beginning of the name. |

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| **Instructions: Copy these six formulas into the appropriate two lists below.**  N2O4  Al2S3  H2O CO CuCO3 C2H6 | |
| **List each ionic formula then write a name using the rules from last week:** | **List each molecular formula, then write a name using the mono/di/tri rules we learned today:** |
| **1)**  **2)** | **3)**  **4)**  **5) 6)** |

**Instructions:**

**A) Circle any substance that is a molecular substance**

**B) Name each molecular substance you circled USING TODAY’S RULES.**

1. CBr4
2. N2P3
3. PCl3
4. MgCl2
5. Hg2O
6. NH3
7. CsBr
8. AgF
9. SnI2
10. Na2O
11. K2S
12. ICl
13. CaBr2
14. BaI2
15. Al2S3
16. N2O
17. GeH4
18. N2Br4
19. P2S5
20. SeO2
21. HgS
22. CuI

**C) Now go back and name each ionic substance USING LAST WEEK’S RULES**

**D) For each molecule below, write the formula and the name**



