

My only big afterschool presence this week is Thursday 340pm to 5pm

I'm here at lunch every day. Come for just 5 minutes, it helps.

Purpose:

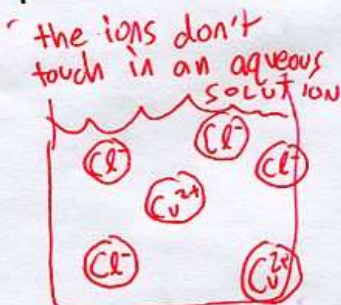
What happened in our 'Nails Lab'?

WARMUP:

Draw an atomic scale picture of each:



a bunch of atoms in an iron nail



a bunch of atoms in $\text{CuCl}_2(\text{aq})$

If possible show live a nail and blue turquoise liquid CuCl_2

#1

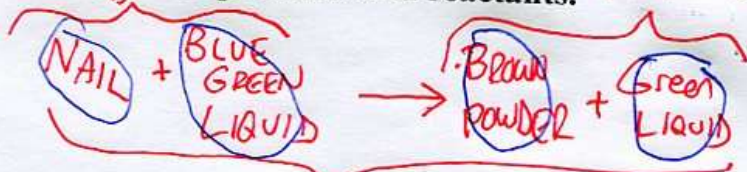
Matter we saw in the lab:

looked like	name	symbol, including phase
turquoise colored water	Copper(II) chloride solution	$\text{CuCl}_2(\text{aq})$
iron nail	solid iron	$\text{Fe}(\text{s})$
copper powder	solid copper	$\text{Cu}(\text{s})$
green liquid	aqueous iron(II) chloride	$\text{FeCl}_2(\text{aq})$

#2

We saw a "chemical reaction"

Reactants with products and reactants. PRODUCTS



A CHEMICAL REACTION



The ratio of copper
to iron is 1:1

In lab today

weigh brown solid ✓

weigh your nails ✓

put your nails back ✓

By the bell

answer Calculations 1, 2, 3

Homework:

Answer all questions well enough to
whiteboard.

CALCULATION #2 is similar to

$$7 \text{ grams Zinc} \times \left(\frac{\quad}{\quad} \right) = \text{moles Zinc}$$

↑ FROM PERIODIC TABLE