

I'm here today and Thursday
3:40pm to 5pm
I'm here at lunch every day.

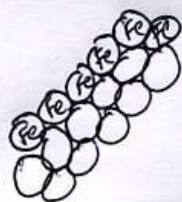
Come for just 5 minutes, it helps.

Purpose:

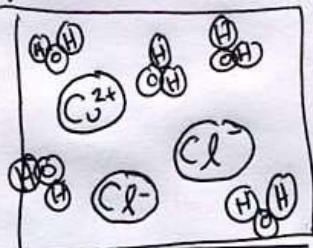
What happened in our 'Nails Lab'?

WARMUP:

Draw an atomic scale picture of each:

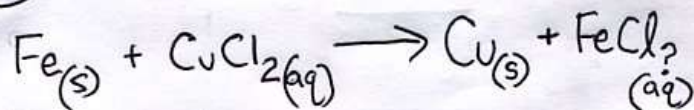


a bunch of atoms in an iron nail



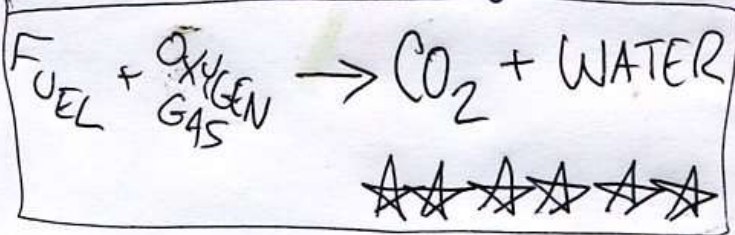
a bunch of atoms in $\text{CuCl}_2(\text{aq})$


#1 What was the Lab reaction?




#2 What do combustion reactions look like?

Question: Write combustion reaction for C_3H_8



 ← 33 grams


 ← 35 grams

SUBTRACT! $35g - 33g = 2 \text{ grams}$

HOW TO
FIND
YOUR
COPPER
MASS

$$2.27 \text{ grams Fe} \times \left(\frac{1 \text{ mole Fe}}{55.85 \text{ grams Fe}} \right) = \text{moles Fe}$$

HOW TO
FIND
MOLES
OF
IRON

 6 grams

 4 grams

How much iron reacted?

$$6g - 4g = 2 \text{ grams}$$

HOW
TO
FIND
IRON
MASS

LAB ADVICE :pa

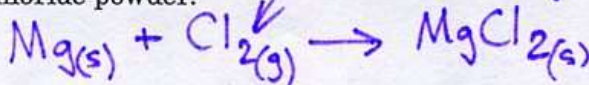
ANSWERS

Name _____

Period _____

Write an equation to describe each. Your answer will only be graded on having the correct formula and the correct phase (solid, liquid, etc). It does not need to obey the Law of conservation of Mass. It does not have to be 'balanced'

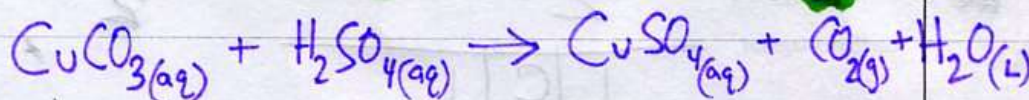
1. Magnesium chunks and chlorine gas react to make magnesium chloride powder.



2. Ethane gas, C₂H₆, reacts with oxygen gas to form carbon dioxide gas and water vapor.

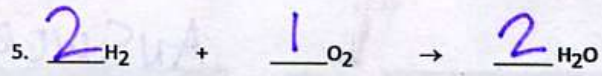


3. Copper(II) carbonate solution and sulfuric acid (H₂SO₄) solution react to make copper(II) sulfate solution, carbon dioxide bubbles and liquid water.

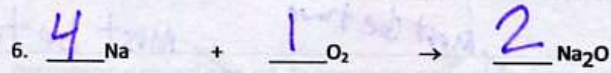


4. Write either the names or formulas (your choice) of all of the REACTANTS shown in the TWO boxes above:

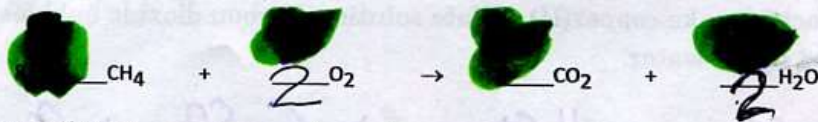
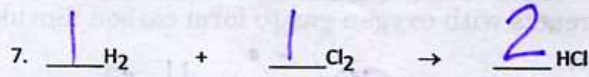
Mg
Cl₂
C₂H₆
O₂
CuCO₃
H₂SO₄



Cartoon Diagram



Cartoon Diagram



Cartoon Diagram

1	C	1
4	H	4
4	O	2+2