

Tonight's homework is Finish "HW 8-5 Adjusting to Reality" This was handed out Monday. It is two pages long, no picture.

Grab a textbook.

Thursday – test

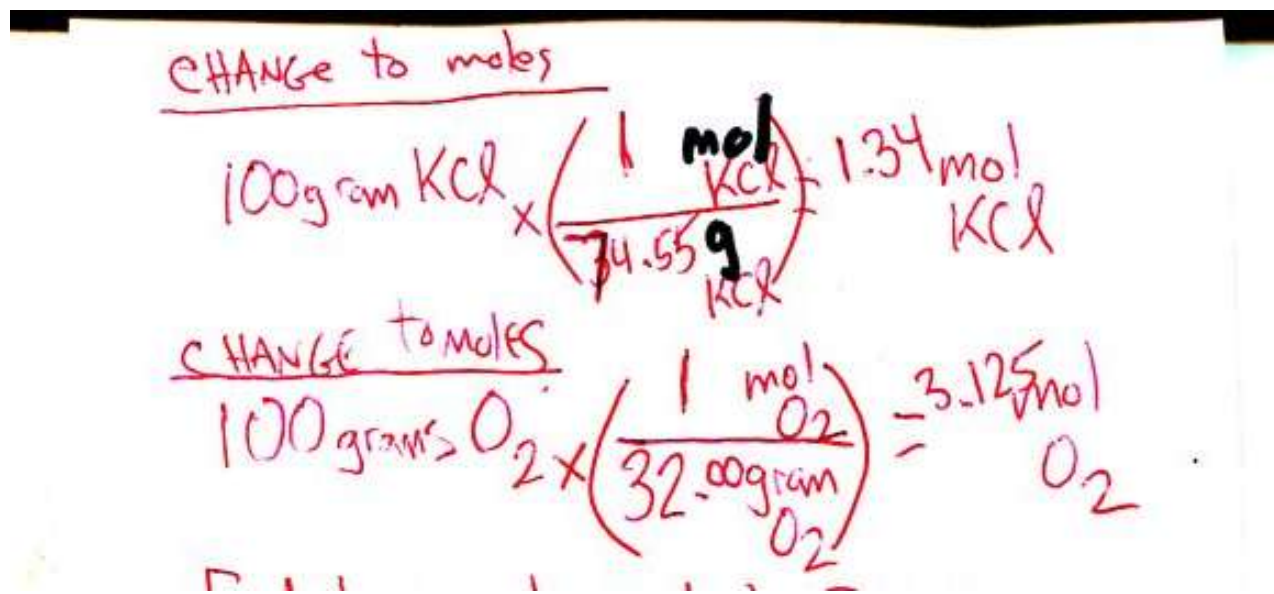
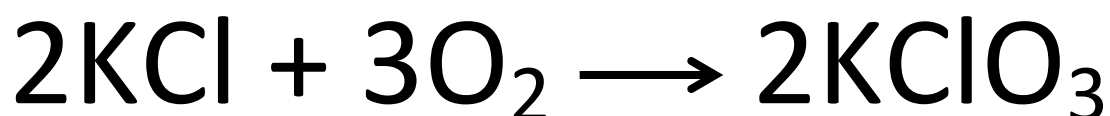
Purpose: (this goes on the right page)

Practice Find which reactant is limiting.

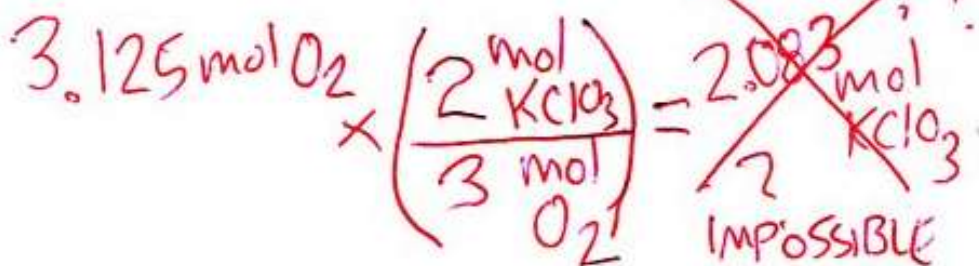
WARMUP (this goes on the left page. just copy, we'll solve together):

Write on the left page of your notebook.

"If you have 100. grams of each reactant, find which is the limiting reactant".



Find how much product forms



The excess reactant in our warmup problem is O₂ (too much)

The limiting reactant is KCl.

(this goes on the right page)

Rule: The excess reactant will be useless for doing stoichiometric calculations.

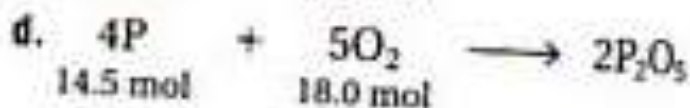
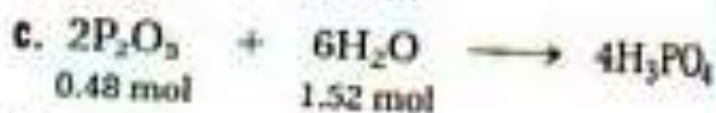
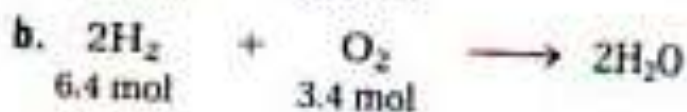
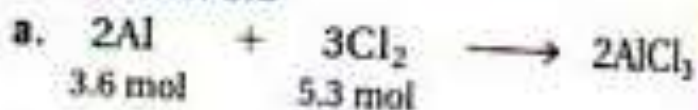
The limiting reactant is what you should use for doing calculations.

Strategy for solving story problems where amounts of more than one reactant are given in the problem: Convert each reactant into moles or grams of *PRODUCT*. The smaller product must be what really happens.

Do the following problems for a stamp and five points classwork credit

(p. 262-263) 44a,b, 45 a,b Blurry scan of the textbook is here ☺

44. For each balanced equation, identify the limiting reagent for the given combination of reactants. 9.3



45. For each reaction in Problem 44, calculate the number of moles of product formed. 9.3

46. For each reaction in Problem 44, calculate

Do any three problems from tonight's homework.

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