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| Review for Test 4. Test 4 is Thursday April 7.  Study everything since the previous stripe on the Website (<http://genest.weebly.com>)  There are two main types of problems. |
| Problem Type 1 |

If 25.0 g of magnesium reacts with excess hydrogen nitrate, how many grams of magnesium nitrate are produced?

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| Review for Test 4. Test 4 is Thursday April 7.  Study everything since the previous stripe on the Website (<http://genest.weebly.com>)  There are two main types of problems. |
| Problem Type 2 |

Determine the mass of carbon dioxide that should be produced in the reaction between 3.74 g of carbon and 11.2 grams of O2. What is the % yield if 11.34 g of CO2 is recovered?

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| Now you try it |
| Problem Type 1 |

If 2.7 grams of silver nitrate react with excess copper metal, how many grams of silver are produced?

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| Now you try it |
| Problem Type 2 |

Suppose 8.61 g of zinc was allowed to react with 1.61 g HCl gas to produce zinc chloride and hydrogen gas.

* 1. What is the balanced equation?
  2. Which reactant is limiting?
  3. According to the grams of ZnCl2 you calculated in (b) for the Limiting Reactant what mass of zinc chloride would be the Theoretical Yield?
  4. Suppose that you actually recovered 1.56 g of zinc chloride. What is your percent yield?