

TODAY

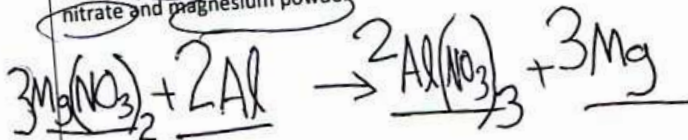
- 1) WARMUP
- 2) Gallery Walk Whiteboarded
- 3) FIRE homework at back.
- 3) Read tomorrow's LAB

Purpose no notebook right page!

Warmup Glue, tape, or staple to your left page.

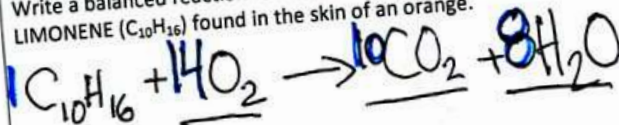
Warmup:

With your periodic table, write a balanced reaction for: magnesium nitrate reacts with aluminum powder to form aluminum nitrate and magnesium powder.



Stop here for now....

Write a balanced reaction for the combustion of 0.003 moles of LIMONENE ($\text{C}_{10}\text{H}_{16}$) found in the skin of an orange.



How many moles of CO_2 should form in the above reaction?

Procedure:

1. Rinse the inside of a clean 100 mL graduated cylinder with a small amount of water. Pour about 50. mL of 0.60M sodium carbonate solution into the graduated cylinder. Record the volume to the nearest 0.5 mL. Pour the solution into a clean 250 mL beaker.
2. Carefully rinse the graduated cylinder two or three times with distilled water. Pour approximately 35 mL of 0.40M calcium chloride solution into the graduated cylinder. Record the volume to the nearest 0.5 mL. Pour the calcium chloride solution into beaker. Describe the resulting reaction in the data table. Stir the contents of the beaker for about 1 minute.
3. With a pencil, write you names and class period on the edge of a piece of filter paper, determine its mass to the nearest 0.01 g, and record the mass in you data table.
4. Set up a funnel holder, funnel, and the filter paper of known mass. Use a 250 mL beaker under the funnel.
5. Wet the filter paper with a small amount of water. Pour the contents of the beaker slowly into the funnel. (Use a stirring rod like you did when decanting.) Be careful as you pour so none of solid flows out of the filter paper or funnel. Use the rubber scraper to transfer as much of the solid from the beaker to the filter paper as possible. Rinse the inside of the beaker with some distilled water to remove more solid. Rinse the beaker two or three times into the funnel.
6. Once all of the solid is on the filter paper and the liquid has all drained through into the beaker, carefully remove the filter paper with the product from the funnel and put the filter paper on a clean watch glass. Dry the filter paper with the product.
7. After the filter paper and calcium carbonate are thoroughly dry, find the mass of the precipitate and filter paper and record it in the data table.

Data Table:

Homework

AND DO #1 and #2

QUIZ covers front page
of this.

We have a test on Friday April 8.

I predict the quiz on Friday April 1 will have a lot of the orange square from today's notes (see above) AND the prelab questions from tonight's homework.