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| 1. Weigh a solid chemical.

I have \_\_\_\_\_\_\_\_ grams of \_\_\_\_\_\_\_\_ | 1. Pour water into any container.

I have \_\_\_\_\_ mL of water. | 1. Calculate.

I have \_\_\_\_\_\_\_\_ moles of the substance from (1) | 1. Calculate the concentration in “M” units
 |

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| 1. Weigh the same solid chemical.

I have \_\_\_\_\_\_\_\_ grams of \_\_\_\_\_\_\_\_\_\_\_ | 1. Pour water into any container.

I have \_\_\_\_\_ mL of water. | 1. Calculate.

I have \_\_\_\_\_\_\_\_ moles of the substance from (1) | 1. Calculate the concentration in “M” units
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| Lab: Make two solutions  Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_of known\* concentration \*We will see next week that our concentration is close but needs one more factor in the calculation. |

Name the two substances used today using correct terminology from our notes from last week.

The solvent is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ The solute is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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| Draw a macroscopic view of the two containers, clearly showing how dark each one is and how much volume is in it | Draw microscopic views of the two containers, clearly showing water and the solute. |
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