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| **If the formula of ‘book’ is C6H12O6, find how many moles of book burned** | **Recall that Cp for water is 4.18 J/g-°C**Q = m cp ∆T**Find how many kilojoules went into the water** |
| **Use ratios to find how much heat would come out from burning 1 mole of book:**$$\frac{your joules}{your moles} = \frac{x}{1 mole}$$ |

**2450. grams**

**of book**

**2050. grams**

**of book**

**200. mL water**

**21.0°C**

**27.0°C**