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| *Solutes + Review a Bit*  CλeMis+ry: http://genest.weebly.com  Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_  Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| 1. What is the (memorized) formula for finding concentration? 2. What is the formula for finding volume if you’re given moles and concentration? 3. In a solution, the substance that is being dissolved is the \_\_\_\_\_.   a. gas b. liquid  c. solute d. solvent   1. From memory, what is the formula of each   sulfuric acid \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  carbonic acid \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  phosphoric acid\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  hydrochloric acid \_\_\_\_\_\_\_\_\_\_\_\_\_  **Identify the solvent and solutes in the following solution.**   1. Dry air contains about 78.1% nitrogen, 21.0% oxygen, 0.9% argon, and trace amounts of other gases.   **solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**  **solute(s):** | 1. Natural gas contains 97% methane (CH4), 1.5% ethane (C2H6), 1% CO2, and 0.5% nitrogen gases.   **solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**  **solute(s):**   1. Rubbing alcohol contains 70% isopropyl alcohol and 30% water.   **solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**  **solute(s):**   1. A standard solution of potassium hydrogen phthalate (KHP), KHC8H4O4 (molar mass = 204.2 g), was prepared by dissolving 5.105 g of KHP in enough water to give 250.0 mL solution. What is the molarity of the KHP solution? (Answer: 0.1000 M) 2. How many grams of solid potassium dichromate, K2Cr2O7 (molar mass = 294.2 g) must be weighed out to prepare 500. mL of 0.200 M K2Cr2O7 solution? (Answer: 29.4 g) |

Cation Anion Formula Name

\_\_\_ Sn2+ + \_\_\_ F– → \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → KC2H3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → \_\_\_\_\_\_\_\_\_ nickel (II) carbonate

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → \_\_\_\_\_\_\_\_\_ magnesium acetate

\_\_\_ Pb2+ + \_\_\_ Cl– → \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → AlCl3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_