|  |  |  |
| --- | --- | --- |
| *Solutes + Review a Bit*CλeMis+ry: http://genest.weebly.com Stop in for help every day at lunch and Tues &Thurs after school! |  | Name\_\_\_\_\_\_\_\_\_\_\_\_\_Period\_\_\_\_\_\_\_\_\_\_\_\_\_ |

|  |  |
| --- | --- |
| 1. What is the (memorized) formula for finding concentration?
2. What is the formula for finding volume if you’re given moles and concentration?
3. In a solution, the substance that is being dissolved is the \_\_\_\_\_.

a. gas b. liquidc. solute d. solvent1. From memory, what is the formula of each

sulfuric acid \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_carbonic acid \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_phosphoric acid\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_hydrochloric acid \_\_\_\_\_\_\_\_\_\_\_\_\_**Identify the solvent and solutes in the following solution.**1. Dry air contains about 78.1% nitrogen, 21.0% oxygen, 0.9% argon, and trace amounts of other gases.

**solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_****solute(s):** | 1. Natural gas contains 97% methane (CH4), 1.5% ethane (C2H6), 1% CO2, and 0.5% nitrogen gases.

**solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_****solute(s):**1. Rubbing alcohol contains 70% isopropyl alcohol and 30% water.

**solvent \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_****solute(s):**1. A standard solution of potassium hydrogen phthalate (KHP), KHC8H4O4 (molar mass = 204.2 g), was prepared by dissolving 5.105 g of KHP in enough water to give 250.0 mL solution. What is the molarity of the KHP solution? (Answer: 0.1000 M)
2. How many grams of solid potassium dichromate, K2Cr2O7 (molar mass = 294.2 g) must be weighed out to prepare 500. mL of 0.200 M K2Cr2O7 solution? (Answer: 29.4 g)
 |

Cation Anion Formula Name

\_\_\_ Sn2+ + \_\_\_ F– → \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → KC2H3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → \_\_\_\_\_\_\_\_\_ nickel (II) carbonate

 \_\_\_\_\_\_ + \_\_\_\_\_\_\_ → \_\_\_\_\_\_\_\_\_ magnesium acetate

\_\_\_ Pb2+ + \_\_\_ Cl– → \_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_ + \_\_\_\_\_\_\_ → AlCl3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_